Non isothermal kinetic study of the aluminium titanate formation in alumina-titania mixtures

(Estudo cinético não isotérmico da formação de titanato de alumínio em misturas de alumina com titânio)

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Abstract

Aluminum titanate (Al₂TiO₅) is a high refractoriness material with excellent thermal shock resistance. Hence it is suitable for several applications at elevated temperatures where insulation and thermal shock resistance are required. Such as components of internal combustion engines, exhaust port liners, metallurgy, and thermal barriers. The thermal instability of Al₂TiO₅ at high temperature is another characteristic of this material that has been studied and controlled by the incorporation of several additives. The Al₂TiO₅ formation from pure oxides presents an endothermic peak in the differential thermal analysis (DTA). The thermodynamic temperature is 1280 ºC. But experimentally, as in every other DTA experiment, these peaks strongly depend on the heating rate: this fact has been extensively employed for the kinetic study of transformation processes and the mechanism determination of chemical reactions. Both activation energies (Eₐ) and nucleation rates can be obtained from these experiments. The present work reports the formation Eₐ of Al₂TiO₅ prepared from pure oxides at air atmosphere by the Kissinger DTA based methods. Previously the particle size distribution of the starting powders together with X-ray diffraction analysis of the starting powders and the resulting materials was carried out. The properties of the Al₂TiO₅ formation were grouped into two groups corresponding to the low and high heating rates, below and over 5 K/min. Eₐ values were obtained after the Avrami (n) constant evidenced that the crystallization mechanism is strongly related to the heating rate, even in the wide range studied which includes the technological ones(0.5-40 K/min).

Keywords: DTA, kinetic, aluminum titanate.

INTRODUCTION

Aluminum titanate (Al₂TiO₅) is a material with high refractoriness and excellent thermal shock resistance. It presents a high melting point (over 2000 K) and anisotropic thermal expansion coefficients [1-4]. Hence it is suitable for several applications at elevated temperatures where insulation and thermal shock resistance are required. Such as components of internal combustion engines, exhaust port liners, metallurgy, and thermal barriers. The thermal
instability of the Al₂TiO₅ is one of its disadvantages [1-4]. Several additives have been employed that demonstrated certain efficiency for stabilizing microstructures [1]. Mullite, talc and zircon are some of the employed additives. The kinetics of the Al₂TiO₅ formation has been investigated isothermally [5, 6]. Depending on temperature and on the reaction progress, different overall kinetic reaction and microstructure evolution have been observed. The non-isothermal kinetic study has not been carried out yet.

The Al₂TiO₅ formation is an endothermic process that occurs at 1280 ºC [1]. Below this temperature it is unstable and dissociates spontaneously. This represents a disadvantage for this material: grain size and stabilizing phase’s incorporation have been the strategies used to overcome this tendency [5, 6].

Differential thermal analysis (DTA) has been extensively employed as a rapid and convenient instrument to study the kinetics of phase transformation processes and chemical reaction mechanisms. This method has been used to investigate crystallisation kinetics in ceramics [7-16], to determine homogeneous crystal nucleation rates and to obtain the activation energy for new phase crystallisation. The present work, reports on the determination of activation energies for the aluminium titanate (Al₂TiO₅) formation from pure oxides at air atmosphere.

**EXPERIMENTAL PROCEDURE**

Commercial starting powders were employed (Alpha-alumina A-16SG, Alcoa Inc., USA) and TiO₂ (Mallinckrodt Chemical Works)), 95% anatase-5% rutile. The mean particle size (D₅₀) of both starting powders is around 1.0 μm.

An initial equimolar mixture was carried out in ethanol. Later the powders were dried at 110 ºC. The differential thermal analyses (DTA) together with the thermogravimetric analysis (TG) were carried out simultaneously for the sintered samples in a Netzsch 409C equipment, using different heating rates φ (from 0.5 to 40 ºC/min) up to 1450 ºC only in the heating cycle. Samples of approximately 200 ± 2 mg were analysed in a platinum crucible and dynamic air atmosphere, using pure alumina as reference material. Original and developed crystalline phases were analysed by X-ray diffraction (XRD, Philips 3020 equipment with Cu-κα radiation, Ni filter, at 40 kV-20 mA). A preliminary crystalline phase analysis as a function of a simple sintering (program) was carried out at air atmosphere between 1200 and 1500 ºC.

Before the kinetic analysis, the DTA peak properties were studied as a function of φ. The resulting ATD peaks presented a Gaussian shape. This kind of curve peak was fitted and the properties of the fitted curves were compared. The obtained centroid was assumed to be the peak temperature (Tₚ). The reacted fraction (Xₐ) was estimated as a relation between the partial integer of the peak at each temperature and the peak area.

**KINETIC ANALYSIS**

A non-isothermal method is experimentally easier than an isothermal method, but the accuracy of the resulting Eₐ value is generally lower than for values determined by isothermal methods [10], for the following reasons: the crystallization peak becomes weaker and broader at slower heating rates, limiting the plotting rang of (1/Eₐ). Under non-isothermal conditions, DTA or DSC curves are used to obtain Eₐ values from peak temperature measurements at different heating rates. Although many mathematical equations have been proposed for the calculations, the Johnson-Mehl-Avrami (JMA) equation and Kissinger method [7-9] are most commonly used. It is important to emphasize that different activation energy values are associated, not only with different crystallization processes, but also with differences in the experimental conditions employed for Eₐ determination [9-10].

**Theory and applied mathematical background**

The Kissinger method, based on Equation A, is commonly used to analyze the kinetic process from DTA, DTG or DSC experiments [10-12].

\[
\ln \left( \frac{\varphi}{T_p^2} \right) = -\frac{E_a}{RT_p} + \ln \left( \frac{AR}{E_a} \right)
\]

(A)

where Tₚ is the maximum temperature of crystallization peak, φ the applied heating rate and R is the universal gas constant. The Kissinger dependence can be used for the direct calculation of activation energy from the changes of maximum of \( \ln \left( \frac{\varphi}{T_p^2} \right) \) vs. temperature dependence with heating rate (Equation A). This plot should give a straight line, which yields the apparent activation energy of process (Eₐ). The detailed theory of the Kissinger approach is described in another publication [7].

The crystallization mechanism can be determined from the shape factor (Avrami constant) n of the endothermic peak using Equation B [9].

\[
n = \frac{2.5}{W_{1/2}} \left( \frac{T_p}{E_a/R} \right)^{2/n}
\]

(B)

where \( W_{1/2} \) is the full width at half maximum of the peak (semi-breath). For lower values of n the crystallization dimension is low, for example a surface crystallization mechanism is supposed rather than volume crystallization. On the other hand, higher values of n are expected only in case of increasing nucleation rates, i.e., >2.5 in diffusion controlled reaction or > 4 in polymorphic transformation.

**RESULTS AND DISCUSSION**

Differential thermal analysis curves for the equimolar alumina-titania (Al₂O₃-TiO₂) mixtures heated at different rates are presented in Fig. 1. Changes in the heating gradients (φ) are accompanied by a change in the shape, the area and the position of the DTA curves. The temperature of
the maximum of the endothermic peak position, $T_p$, shifts to higher temperature from 1350 ºC to 1420 ºC, as the heating rate increases from 0.5 to 40 K/min.

Table I shows the different properties of the evaluated peaks at different rates. Although the thermodynamically formation temperature is 1553 K [5, 6], in the present non-isothermal experiments (constant heating rate), the endothermic peak was only detected at 0.5 K/min at 1622 K. During a lower rate experiment the endothermic peak was not detected. This can be explained by the small free energy formation (driving force). The driving force of the formation is small ($\Delta G^0_{\text{Al}_2\text{TiO}_5} < 1500$ J/mol) for temperatures below 1422 ºC and remains below 2500 J/mol for temperatures under 1500 ºC, temperatures commonly used to processes this kind of material [5].

In the studied conditions, the aluminium titanate formation is highly influenced by the heating rate. The formation temperature was ~ 1700 K for the 40 K/min experiment, which is almost 150 degrees over the equilibrium temperature. In addition, $T_p$ showed a 70 degrees shift in the rate range studied. This fact encourages the application of the non-isothermal kinetic analysis carried out in the present work.

![Figure 1: DTA curves, endothermic peak corresponding to the Al$_2$TiO$_5$ formation from pure oxides.](image1)

![Figure 2: DTA peak height, width and area as a function of $T_p$.](image2)

Table I - DTA peak properties.

<table>
<thead>
<tr>
<th>$\phi$ (K/min)</th>
<th>$T_p$ (ºC)</th>
<th>$T_p$ (K)</th>
<th>Area</th>
<th>Width</th>
<th>Height</th>
<th>Gaussian fitting result ($R^2$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Formation equilibrium temperature [5, 6]</td>
<td>1280</td>
<td>1553</td>
<td>-</td>
<td>-</td>
<td>-</td>
<td>-</td>
</tr>
<tr>
<td>0.5</td>
<td>1349</td>
<td>1622</td>
<td>11.9</td>
<td>8.57</td>
<td>-1.11</td>
<td>0.88</td>
</tr>
<tr>
<td>1</td>
<td>1365</td>
<td>1638</td>
<td>28.1</td>
<td>9.87</td>
<td>-2.27</td>
<td>0.81</td>
</tr>
<tr>
<td>1.5</td>
<td>1368</td>
<td>1641</td>
<td>44.0</td>
<td>11.69</td>
<td>-3.06</td>
<td>0.95</td>
</tr>
<tr>
<td>2.5</td>
<td>1375</td>
<td>1648</td>
<td>53.7</td>
<td>11.34</td>
<td>-3.77</td>
<td>0.93</td>
</tr>
<tr>
<td>5</td>
<td>1379</td>
<td>1652</td>
<td>120.9</td>
<td>12.97</td>
<td>-7.44</td>
<td>0.98</td>
</tr>
<tr>
<td>10</td>
<td>1387</td>
<td>1660</td>
<td>282.4</td>
<td>17.49</td>
<td>-12.8</td>
<td>0.993</td>
</tr>
<tr>
<td>20</td>
<td>1402.5</td>
<td>1675.5</td>
<td>535.4</td>
<td>22.8</td>
<td>-18.73</td>
<td>0.991</td>
</tr>
<tr>
<td>40</td>
<td>1421</td>
<td>1694</td>
<td>818.9</td>
<td>26.18</td>
<td>-24.95</td>
<td>0.994</td>
</tr>
</tbody>
</table>
Fig. 2 shows peak height, width and area. Two zones can be clearly defined as a function of $T_p$. This fact could be an evidence of a change in the detected mechanism. While for the low rate (0.5-2.5 K/min) the increment is small, this is more abrupt for experiments with higher heating rates (5-40 K/min).

**Kinetic results**

Fig. 3 shows the Kissinger plots ($\ln(\phi/T_p^2)$ versus $T^{-1}$) in the whole $\phi$ range. The apparent activation energy values are shown in Table II. On the other side, Figs. 4 and 5, respectively show the Kissinger plots of the two heating rate ranges grouped in the previous section.

### Table II - Apparent activation energies ($E_a$).

<table>
<thead>
<tr>
<th>Method</th>
<th>Kissinger Apparent activation energies $E_a$ (kJ/mol)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Complete rate 0.5-40 (K/min)</td>
<td>1480</td>
</tr>
<tr>
<td>Low rate range 0.5-2.5 (K/min)</td>
<td>1313</td>
</tr>
<tr>
<td>High rate range 5-40 (K/min)</td>
<td>1081</td>
</tr>
</tbody>
</table>

The apparent activation energies are shown in Table II. The obtained values are expectably high because the formation temperature is high (1280 °C). It is valid to note also that the values of the enthalpy of formation of are of the same order [17-18]. In a former the isothermal kinetic study [5, 6], no activation energy was estimated, hence the present is the first reported activation energy for this compound. The same grouping was carried out between low and high heating rate. An important difference can be observed: the activation energy at low rates is 22% higher than the ones observed at higher heating rates, showing that the activation energy, calculated from Fig. 3, lacks representativeness. However, the other two values have a greater validity.

The aluminium titanate conversion as a function of the temperature for the different heating rates is shown in Fig. 6. In every case the sigmoidal shape can be observed, where conversion rate is low at the beginning and at the end of the process but fast in the intermediate stages. The low initial rate is explained by the time necessary for a significant number of particles to form nuclei of the new phase sufficiently large and stable to be able to grow. During the intermediate phase, transformation is fast: nuclei, which are numerous and of sufficient size, start to grow, consuming the old phase, while the new nuclei continue to form in the old phase. However, when the transformation is about to conclude, there is very little non-transformed phase nuclei that can form or grow at its expenses. This causes the new phase growth to slow down. The Kissinger theory offers the mathematical formalism to describe this kind of process [7].

The sigmoid form is evident in Fig. 6. Linear fitting...
was carried out between 10 and 90% conversion (X) in the double logarithmic plot (Fig. 7). From this, the kinetics values can be obtained by the JMA method [10]. This was carried out and the results are similar to the ones obtained by the Kissinger method, not shown in the present work.

The Avrami constant brings information about the formation mechanism, but this information is ambiguous and not clear, although some authors propose that the value corresponds to the dimensionality of the crystalline products. Several systems present non-integer values or values over 3, showing the dimensionality hypothesis is not robust [12]. However, it tends to be used to establish whether the reaction mechanism is one or more than one [12]. The Avrami constants for each heating rate, calculated with Equation B and the corresponding activation energy (Table II), are shown in Table III.

First of all, it is important to note that the n constant values obtained are not constant, showing there is not an unambiguous mechanism for the aluminium titanate formation from oxides. In all the cases the values are between 2 and 4.8. This fact evidences an important difference in the mechanism, depending on the heating range. From the grouping performed (low and high heating rate ranges), important differences were observed as well: while for low heating rates a high (around 4) value is obtained, for the higher heating rates the value is low (between 3.3 and 2.1). Moreover, within the grouped ranges, the value of n also decreases with φ (logarithmic plot). This can be easily observed in Fig. 8.

These results are in a way concordant with the results obtained by the isothermal kinetic analysis performed by other authors [5, 6]. In these articles two reaction stages were identified during isothermal treatment at ~ 1580 K, when the free energy of formation is small. The initial one is characterized by the nucleation and growth of Al₂TiO₅ cells in a virtually unreacted Al₂O₃-TiO₂ matrix. The concentration of nucleation sites appears to be temperature independent and it is thought to be related to residual impurities. Unreacted TiO₂ also is found to provide an effective diffusion path.
for the rapid transport of material during cell growth. The diffusion controlled final reaction stage corresponds to the elimination of Al₂O₃ and TiO₂ dispersoids trapped during the initial growth of Al₂TiO₅ cells.

The only confirmed conclusion in view of the non-isothermal results is that the mechanism varies with the heating rate. Perhaps more than one reaction occurs (for example consecutive) and the dependence with the heating rate of these is dissimilar and thus why two or more ranges can be observed by this method. In the high heating rate range perhaps the first stage is completely instantaneous and the evaluated $E_a$ and $n$ correspond to the second stage described by these authors, and the low heating rate range values correspond to the first stage control. In other words, they are controlled by the nucleation and growth rates, which in turn depend on the diffusion rate of the products to the nuclei of growth, and probably these are affected by the conversion to Al₂TiO₅.

CONCLUSIONS

A non-isothermal kinetic study of the formation of aluminum titanate was carried out by differential thermal analysis of pure oxide mixtures of D₅₀=1.0 μm. Although the Al₂TiO₅ formation temperature is 1280 °C, under the worked conditions the reaction only occurred at higher temperatures (1340-1420 °C) and at the same time. It is strongly correlated with the heating rate. In all cases, symmetrical and Gaussian peaks were found. Two zones were identified with different behavior below and over 5 K/min from the peak properties. The formation energy $E_a$ could be obtained by the method of Kissinger from the heating rate dependence, with consistent values. Two zones could be defined with different behavior, depending on the heating rate: below 5 K/min, the $E_a$ value was 1313 KJ/mol, and above that value $E_a$ is ~25% lower (1081 KJ/mol). The Avrami constant $n$ was evaluated in the whole heating rate. These bring some information on the formation mechanisms. At least two mechanisms were found, depending on the heating rate. A high $n$ (between 3 and 4) was calculated for low heating rates and values between 2 and 3 were found for high heating rates for the $n$ constant.

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REFERENCES