A New Triazene Ligand Immobilized on Triacetylcellulose Membrane for Selective Determination of Mercury Ion

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Este trabalho apresenta um novo sensor químico óptico, simples e de baixo custo, para a determinação de íon mercúrio (II) em solução aquosa. A membrana sensível a mercúrio foi preparada pela incorporação de triazina (((E)-1-(2-etoxifenil)-3-(4-nitrofenil) triazol-1-ene) como um ligante adequado para Hg(II) sobre uma membrana de triacetilcelulose. O estudo espectrofotométrico do complexo formado entre o ligante triazene (L) e ions Hg²⁺ em solução de dimetilformamida (DMF) mostrou uma grande constante de estabilidade para o íon mercúrio complexado. A membrana responde ao íon mercúrio pela troca reversível da cor de laranja para verde em solução tamponada em pH 3.0. Uma relação linear foi observada entre a absorbância da membrana em 405 nm, na faixa de 7 a 90 µg mL⁻¹, com um limite de detecção de 64 ng mL⁻¹ em soluções aquosas, em pH 3.0. O sensor óptico foi aplicado com sucesso na determinação de mercúrio em amostras de água enriquecidas com o analito.

Keywords: optical sensor, Hg²⁺ ion, triazene (((E)-1-(2-ethoxyphenyl)-3-(4-nitrophenyl) triazol-1-ene) ligand

Introduction

Mercury is one of the most toxic elements in the world and represents a major toxicity to microorganisms and environment even in low concentrations.¹ Accurate and sensitive determination of mercury is now a challenging task. Inorganic mercury has been reported to produce harmful effects at concentrations as low as 5 µg L⁻¹.¹ Inorganic mercury may be converted to methyl mercury in marine environment that is even more toxic than inorganic mercury.³ Thus, the development of simple methods for selective determination of mercury in trace amounts in different matrices is critical. A wide range of analytical methods has been used for the determination of mercury in real samples. These include spectrophotometry,⁴ neutron activation analysis,³ anodic stripping voltammetry,⁶ X-ray fluorescence spectrometry,⁷ inductively coupled plasma mass spectrometry,⁸ electro-thermal atomic absorption spectrometry,⁹ fluorescence spectrometry,¹⁰ and cold vapor atomic absorption spectrometry.¹¹ However, most of these techniques are rather expensive and time consuming and may not be available in all laboratories. A few works have been introduced on the development of optodes which are based on spectrophotometric method for Hg(II).¹²⁻¹⁶ Meanwhile, during the past decade, the development of optical chemo-sensors for the determination of heavy metal ions has become a rapidly expanding area of analytical chemistry because it offers certain advantages.

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over electrochemical sensors. Optical chemical sensors (optodes) are considered to be better studied for monitoring of toxic ions/species. Optodes are easy to prepare, selective for pre-concentration of target analyte in the sensing matrix, and there is no need to separate reference devices. Thus, in recent years, a number of optodes with different optical principles have been reported for the determination of low levels of mercury. The current paper encompasses the fabrication of an optode for determination of mercury in aqueous solutions.

**Experimental**

**Material and instruments**

All of the chemicals used were of analytical reagent grade (Merck, Darmstadt, Germany). De-ionized water (Abpakhsh, Tehran, Iran) was used throughout. A 1000 µg mL⁻¹ Hg(II) stock solution was prepared by dissolving 0.1708 g Hg(NO₃)₂·H₂O in de-ionized water and 0.5 mL of 1 mol L⁻¹ HNO₃ and diluted to 100 mL in a volumetric flask. Sample solutions were prepared by appropriate dilution. The triazene ligand was synthesized and its structure with Hg²⁺ ion has been confirmed (Figure 1). The 1.166 × 10⁻³ mol L⁻¹ triazene ligand solution was prepared by dissolving 10.0 mg ligand L in 30.0 mL ethylenediamine. A Shimadzu 1601 (Kyoto, Japan) PC UV-Vis spectrophotometer with 1 cm cell quartz was used for recording all spectra and absorbance measurements. A Jenway 3510 pH-meter (Staffordishe, UK) with a combined glass electrode was used after calibration against standard Merck buffer for pH adjustment.

Figure 1. The molecular structure of [Hg(C₁₅H₁₃N₄O₃)₂].

**Determination of formation constant**

In order to determine stoichiometry and stability of the resulting metal-ligand complexes, in a typical procedure, 2.0 mL of 5.0 × 10⁻⁵ mol L⁻¹ ligand solution in dimethyl formamide (DMF) were placed in the spectrophotometer cell and the absorbance of solution in range of 315-600 nm was measured. Then a known amount of the concentrated solution of Hg²⁺ in 1.3 × 10⁻³ mol L⁻¹ DMF was added in a stepwise manner using a 10 µL Hamilton syringe. The absorbance spectrum of the solution was recorded after each addition. The absorbance curves are shown in Figure 2. By addition of Hg²⁺, a decrease in absorbance is observed at 405 nm. By decreasing the absorbance at 405 nm, a new peak at about 465 nm is formed that corresponds to the formation of an Hg²⁺-L complex. The resulting plots of the absorbance at 465 nm against metal ion/ligand mole ratios are shown in Figure 2. From the sharp inflection point observed for Hg²⁺ at a mole ratio of 2, it can be concluded that a 2:1 complex of [L₂-Hg] is formed in DMF solution. The formation constant of the resulting complex between Hg²⁺ and L ligand was evaluated to be equal to 7.81, from the absorbance versus [Hg²⁺]/[L] mole ratio data using known equations and utilizing a non-linear least-squares curve fitting program, KINFIT.
Preparation of the sensor membrane and measurement

For immobilization of the triazene ligand L, triacetylcellulose membrane was treated with a $1.166 \times 10^{-3}$ mol L$^{-1}$ ligand solution, in ethylenediamine for 2-5 min at ambient temperature. The resulting orange color membranes were thoroughly washed with detergent solution and water. Prepared membranes were kept under water when not in use. Each time six membranes were made and they can still be used after 1 month when kept under water. A 1 cm $\times$ 3 cm piece of the fabricated membrane sensor was cut and placed inside the quartz cell of the spectrophotometer. All the measurements on the triacetylcellulose membranes were performed in aqueous medium. Tap water samples were collected from the laboratory water tap at Payame Noor Shiraz University, Iran. The tap water samples were spiked with different amounts of Hg$^{2+}$ and pH adjusted to 3.0 before analysis. For analysis, about 2.5 mL of the samples were transferred to a 1 cm quartz cell equipped with the membrane sensor. The absorbance was then measured at 405 nm and subtracted from an absorbance reading for a buffer solution at the same wavelength. The Hg$^{2+}$ concentration was then derived using an ordinary calibration curve method.

Results and discussion

The first extensive investigation of the coordination chemistry of a triazene derivative (1,3-diphenyltriazene) was carried out in 1887 by Meldola and Streatfield. The interaction of triazene derivatives with mercury ion has been studied in the past few years by several authors. The affinity of the ligand L toward Hg$^{2+}$ must be related to the coordinate interaction between ligand L and Hg$^{2+}$ ion. In a ligand molecule, the O and N atoms can play the role of an electron donor which can coordinate metal ions as electron acceptors. On the other hand, on the basis of Pearson’s hard-soft [Lewis] acid-base principle, Hg$^{2+}$ is a soft acid and triazene ligand is a soft Lewis base. Organic ligands containing unsaturated nitrogen atoms can be regarded as soft Lewis bases. As such they may reveal high tendency to form stable coordination complexes with numerous transition metal ions, particularly those that can be regarded as soft Lewis acids [e.g., Cu(I), Ag(I) and Hg(II)]. Therefore, it was also known that certain ligands formed the most stable complexes with Hg$^{2+}$. Soft Lewis acid binds to soft Lewis base to give covalent complexes. These interactions are dominated by the energies of the participating frontier molecular orbitals (FMO).

Spectral characteristics

Figure 3 shows the absorption spectra of immobilized triazene ligand, which was obtained after being equilibrated in pH 3.0 buffer solution containing different concentrations of mercury (0.0-5.0 µg mL$^{-1}$ Hg$^{2+}$, by addition of 0.5 µg mL$^{-1}$ Hg$^{2+}$ at each interval). The spectral change (increase in absorption band at 433 nm and decrease in absorption band at 405 nm) is the result of increase of mercury ion concentration in the membrane and complex formation. The maximum absorbance of the immobilized ligand L is located at 405 nm. The wavelength of 405 nm was selected for further studies because of higher selectivity and sensitivity at this wavelength.

Effect of pH of test solution on the sensor response

The influence of test solution pH on the response of the proposed Hg$^{2+}$ ion-selective optical sensor is illustrated in Figure 4. The response characteristic of the membrane sensor was highly dependent on pH. The absorbance measurements were made for 10 µg mL$^{-1}$ mercury ion in the pH range of 1-8 at 405 nm. As it is obvious from Figure 4, the absorbance
increased rapidly by changing the pH from 2.0 to 3.0, while it decreased at pH values higher than 4.0. This phenomenon might be due to the fact that at lower pH values (pH < 3.0), complexation is weak. On the other hand, the reduced optical response of the proposed sensor at pH > 3.0 could be due to a possible hydroxide formation of mercury ions. Thus, in subsequent experiments, pH 3.0 was used for further studies.

Sensor response time

Figure 5 shows the profile of the response of Hg$^{2+}$ optical sensor at 405 nm with time. The response time of the present optode was tested by recording the absorbance change at 405 nm from a pure pH 3.0 buffer to a buffered mercury solution of 10 µg mL$^{-1}$ Hg$^{2+}$. It can be seen that the time taken to achieve to the 95% steady state response is within 15 min for 10 µg mL$^{-1}$ Hg$^{2+}$ at pH = 3.0.

Dynamic range

Figure 6 shows the absorbance signals of the optode film to various concentrations of Hg$^{2+}$ ions in the range of 7.0-90.0 µg mL$^{-1}$ at pH 3.0 and 90.0 µg mL$^{-1}$ was found as the concentration of Hg$^{2+}$ ion that saturates the film. The regression equation was:

$$\Delta A = 0.0003 \times C_{\text{Hg}^{2+}} + 0.0609$$  \hspace{1cm} (1)

with a correlation coefficient of 0.9998, where $\Delta A$ is the increase in absorbance of the film at 405 nm for a fixed time of 15 min, and C is the concentration of Hg$^{2+}$ in µg mL$^{-1}$. The detection limit which is based on 3σ of the blank was calculated to be 64.0 ng mL$^{-1}$.

Regeneration of the optode membrane

Some reagents, including HCl, NaOH, H$_2$SO$_4$, HNO$_3$, and ethylenediamine, were studied as regenerating reagents. It was found that the best result was obtained by applying ethylenediamine which gave short membrane regeneration times (10-25 s). After this regeneration and for the next mercury concentration measurement, the optode must be placed in buffer pH 3.0 for 10-15 min.

Selectivity

The major property of the optode membrane, its selectivity, reflects its relative response to the analyte ion and to the other ions present in solution. Thus, the influence of a number of common metal ions on the absorbance of the proposed Hg$^{2+}$ optical sensor was carried out. To determine of selectivity of the optode membrane, the membrane was tested for the determination of 10.0 µg mL$^{-1}$ of Hg$^{2+}$ ions in the presence of some metal ions, including Cd$^{2+}$, Cu$^{2+}$, Fe$^{3+}$, Pb$^{2+}$, Zn$^{2+}$, Ni$^{2+}$, Co$^{2+}$, and Fe$^{3+}$. The species were considered as interference if they caused an analytical variation of 5% or more when compared to the analytical signal obtained in the absence of the interfering species. At the applied pH value, no interference was observed from even 50-fold excess of the interfering ions.

Accuracy and analytical applications

The proposed optical sensor was found to work well under laboratory conditions. In order to test the practical application of the present sensor, the distilled water and tap water samples spiked with different amounts of mercury ions were measured by the proposed optode. The mercury content of water was analyzed by standard addition method and then determined by the proposed optode (Table 1). From the data given in Table 1, it is rapidly apparent that the present optode is useful for the determination of mercury in real samples.
Table 1. Result of mercury (II) ion determination in spiked samples

<table>
<thead>
<tr>
<th>Sample</th>
<th>Mercury(II) added / (µg mL⁻¹)</th>
<th>Mercury(II) found / (µg mL⁻¹)</th>
<th>RSD / %</th>
<th>Recovery / %</th>
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<td>10</td>
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<td>0.097</td>
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<tr>
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</table>

*Average of 3 determination found by proposed method; not detected. RSD: relative standard deviation.

Conclusions

In this work, a new triazene ligand (E)-1-(2-ethoxyphenyl)-3-(4-nitrophenyl) triaze-1-ene was considered as a suitable complexing agent for construction of a triacetylcellulose optical membrane selective sensor for mercury(II) ion detection at low concentration levels. On the basis of the results presented in this work, the proposed Hg(II) ion-selective optode has many advantages including: easy preparation, low cost, fast response time, wide dynamic range, low detection limit, and good reproducibility. It was applied in order to determine the concentration of mercury(II) ions in water samples.

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References


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