

Article

Passivation of Iron in Solutions Containing Borate Using Rotating Ring-Disc Measurements

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Investigou-se o comportamento eletroquímico do ferro em soluções neutras de perclorato de sódio contendo pequenas concentrações de borato de sódio, através da técnica de voltametria cíclica e empregando-se eletrodo rotatório de disco e anel. Nas soluções de perclorato, a passivação do ferro ocorreu após o pico anódico correspondente à zona ativa, no qual foi observada a dissolução de espécies Fe^{2+} e Fe^{3+} . Quando é adicionado borato à solução, tanto a densidade de corrente do pico anódico quanto a quantidade de espécies ativas de ferro solúveis diminuem. Para uma concentração constante de perclorato e borato, o aumento do pH do meio mostrou ser menos efetivo na promoção da passivação do ferro do que a presença de borato. Os resultados demonstraram que a interação entre o íon borato e a superfície eletródica ocorre por adsorção química, de acordo com a isoterma de Langmuir.

The electrochemical behavior of iron in neutral solutions of perchlorate and low borate concentrations was investigated using cyclic voltammetry and the rotating ring-disc electrode techniques. In solutions containing just perchlorate, iron passivation was observed at potentials higher than the potential of an anodic peak corresponding to the active range at which the dissolution of Fe^{2+} and Fe^{3+} species was detected. When borate is added to the solution, this anodic peak decreased as well as the amount of the active soluble species. The increase of pH, keeping the borate and perchlorate concentrations constant, was found to be less effective at promoting iron passivation than the increase of borate concentration. The results demonstrated that the interaction between borate ion and the electrode surface occurs by chemical adsorption according to the Langmuir isotherm.

Keywords: *iron passivity, borate passivity, anodic films*

Introduction

During the last 30 years, the buffer sodium borate-boric acid has been intensively used as an electrolyte in many electrochemical studies that investigate the behavior of several metals and alloys, mainly iron, nickel, copper, cobalt and steel¹⁻¹⁰. The concentrations of each component of this buffer are always high *e.g.* 0.1 mol/L or higher, in order to promote a low ohmic drop in the solution. In these studies the borate buffer acts as support electrolyte or as a blank test to which the results of experiments using aggressive anions, such as chloride and sulphate, could be compared. The borate buffer solution is often used in these

studies because in this electrolyte the cathodic reduction of initial oxide formed in air is easy¹¹ and the anodic passivation is very efficient¹¹⁻¹⁸.

Nevertheless most authors have not investigated the influence of the borate anion on the studied metal dissolution and passivation processes. Some authors have considered this anion to be a corrosion inhibitor, mainly for iron and steel, and have explained its performance as a pH controller buffer¹¹⁻¹⁸. Only a few authors have considered the borate anion as an adsorption inhibitor¹²⁻¹⁵ and studies in which this anion is used in small concentrations are very scarce¹²⁻¹³. This kind of study should be carried out when an adsorption inhibitor is investigated. No data in the litera-

ture was found concerning the active species dissolution in the presence of borate.

The use of a rotating disc electrode presents several advantages because there is an important feature that distinguishes it from the other reaction surfaces in a flowing fluid, that is, the thickness of the hydrodynamic boundary layer is constant along the whole disc surface. This fact assures that the mass transport to any point of the disc surface is identical. Such a reaction surface is uniformly accessible from the standpoint of diffusion¹⁹.

Therefore, in experiments using RRDE, any soluble species produced at the disc, that is electrochemically active, could be homogeneously collected by the Pt ring and therefore identified by this way.

The equations for convective diffusion, derived from the general equation of Navier-Stokes, are simplified when the reaction site is the surface of a rotating disc. Such simplifications make the use of RDE very convenient in electrochemical experiments.

Nevertheless, the equations derived for RDE could not be applied for very small or very large values for the electrode rotation rate (ω). When ω is smaller than 100 rpm, the hydrodynamic boundary layer becomes large and approaches the disc radius. Under this condition, the derived equations are not followed. On the other hand, for values of ω larger than 10,000 rpm, the flow regime becomes turbulent. Values of the Reynolds number greater than 2×10^5 are not convenient for RDE experiments because this represents the limit for laminar flux conditions. This fact is important because when the flux becomes turbulent the equations of mass transport are also no longer valid.

For these reasons the range of ω generally used in RDE studies is between 100 and 10,000 rpm²¹.

The aim of this study is to compare the electrochemical behavior of iron in a solution of sodium perchlorate 0.1 mol/L containing or not small amounts of borate in the absence of its acid pair, as in this manner the borate buffering capacity is decreased. At these conditions the use of rotating ring-disc electrode is used in order to control the dissolution of active species before and after of borate addition.

Experimental

The working electrode consisted of a Gallard-Schlesinger 99.999% pure iron disc embedded in PTFE, with an exposed area of 0.1886 cm². This electrode was polished on SiC paper down to grade 1000 and 0.1 μ m alumina paste, and then washed, rinsed with bidistilled water and ethanol and dried. A potential of -1600 mV in relation of a saturated calomel electrode (SCE) was applied for 10 min in the same working solution for reduction of air-formed films. For the rotatory ring-disc electrode (RRDE) experiments, a plati-

num ring with an area of 0.0651 cm² with a collection factor of 0.26²¹ has been used. All experiments have been run with the electrode rotating at 1000 rpm. The solutions used were 0.1 mol/L sodium perchlorate to which 0 to 3000 ppm borax had been added. Whenever necessary, the pH was adjusted by adding a solution of 0.1 mol/L NaOH. All chemicals were analytical grade. The equipment consisted of a Pine bipotentiostat Model AFRDE 5, a BBC Model SE 780 X-Y recorder and a Pine analytical rotor. The electrolytical cell consisted of a pyrex glass flask with four entrances. Potential has been measured against a saturated calomel electrode (SCE) provided with a double sintered glass junction and a Luggin capillary. The potentials are referred to it throughout the text. The counterelectrode was a platinum wire. The experiments have been run at room temperature and under a N₂ atmosphere.

Results and Discussion

Voltammetric experiments with RRDE in 0.1 mol/L sodium perchlorate with the pH adjusted to 9 were carried out. Figure 1.a shows the results of disc current i_d for the two first cycles obtained at a scan rate of 50 mV s⁻¹. The ring current detected along the potential scan of the disc under oxidizing (+600 mV) or reducing conditions (-800 mV) is also shown. The voltammogram in perchlorate solution is characterized by an extensive active region with one single current peak (P_A), after which the electrode passivates. Along both the active as well as the passive zones iron dissolution occurs as Fe²⁺ species as is shown by peaks 1, 2 and 3 on the ring curve (Fig. 1.b). Dissolution by formation of Fe³⁺ species is also observed during the oxidation of the disc electrode (Fig. 1.c). The total anodic charge evaluated by integrating the i_d curve referring to the first cycle is 79.6 mC cm⁻². The charge under the i_r ring curves, kept in oxidizing conditions as well as in reducing conditions during the positive going scan allows calculation of the total charge density involved in the dissolution processes by both Fe²⁺ and Fe³⁺ species, as equal to 4.3 mC cm⁻². This value corresponds to 5.4% of the total anodic charge of the disc. Thus, the anodic peak presented on the disc voltammogram corresponds to the formation of a film containing Fe(II) species, with some simultaneous transport of Fe²⁺ species to the solution. As the potential increases, part of the Fe(II) of the film oxidizes to Fe(III), with simultaneous dissolution as Fe²⁺ and Fe³⁺ species. The formation of these Fe(III) species on the surface leads to passivation, while the film forms at the metal / film interface and partially dissolves at the film / solution interface. These facts indicate that Fe(II) species are constantly produced during all the anodic region of potentials. Inverting the potential scan a significant hysteresis confirms the existence of the passive film, which suffers a partial reduction represented by a cathodic peak

of small intensity (P_C). After one complete voltammetric cycle the total cathodic charge is smaller than the anodic one, even taking into account the dissolutive processes. This fact shows that the film reduction occurs only partially or otherwise at potentials as negative as that of H_2 evolution. During the reduction process, Fe^{2+} species are caught on the ring (peak 4, Fig. 1.b). Probably part of the $Fe(III)$ species of the film are reductively dissolved to Fe^{2+} species.

With continuous potential cycling and also increasing scan rate, both the anodic as well as the cathodic peak increase and separate from each other. Thus the processes represented by peaks P_A and P_C correspond to a redox pair, which becomes more irreversible with cycling or with increasing potential scan rate. The growing distance between anodic and cathodic peaks may be related with the ageing processes of the film²², which dehydrates, getting less conductive and rendering its reduction and further reoxidation more difficult. The increase of the intensity of these peaks with the progressive cycling is a probable consequence of increasing the active area of the electrode, due to higher rugosity of the surface, resulting from the dissolution processes.

Thus iron passivation in perchlorate solutions is not an efficient process because it requires the transport of a great amount of charge. Perchlorate anion is not expected to adsorb on the metallic surface and thus the film-forming mechanism does not occur by a topochemical pathway, but probably by precipitation, requiring the species saturation near the electrode, with high current transport.

Different amounts of borate, from 500 to 3000 ppm were added to the 0.1 mol/L sodium perchlorate solution and Fig. 2 shows the obtained voltammograms in the same

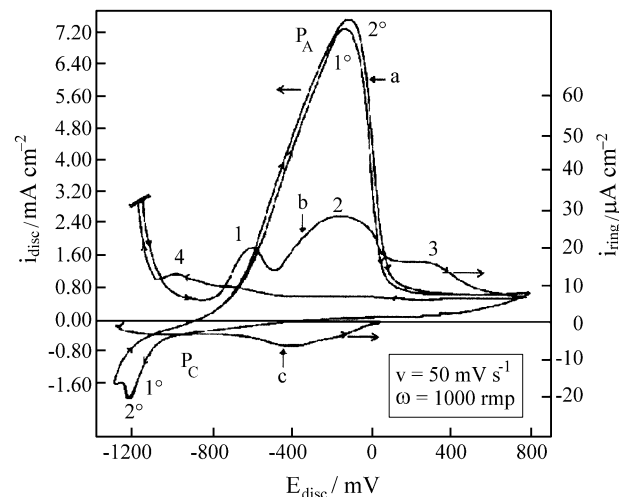


Figure 1. a - Cyclic voltammograms of iron in 0.1 mol/L sodium perchlorate (pH). First and second cycles. b - Ring current density ($E_{RING} = 600 mV$) during disc potential scan, showing four peaks (1, 2, 3 and 4). c - Ring current density ($E_{RING} = -800 mV$) during disc potential scan.

condition as those in Fig. 1. The anodic peak P_A as well as the cathodic peak P_C diminish according to borate concentration and their potentials are shifted to more positive values. In the inverse scan, the curves cross the abscissa axis always at about the same potential ($-400 mV$), which may indicate that the same species is reduced under all conditions.

The influence of consecutive potential cycling on the voltammetric profiles has been verified for borate concentrations from 1000 to 3000 ppm. Figure 3 presents the results obtained for the first five cycles in the perchlorate

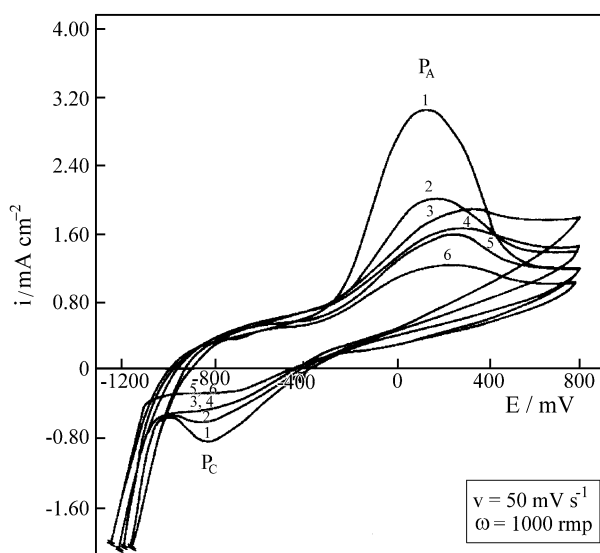


Figure 2. Cyclic voltammograms of iron in 0.1 mol/L sodium perchlorate (pH 9) with the addition of borax: 1 - 500 ppm; 2 - 1000 ppm; 3 - 1500 ppm; 4 - 2000 ppm; 5 - 2500 ppm; 6 - 3000 ppm.

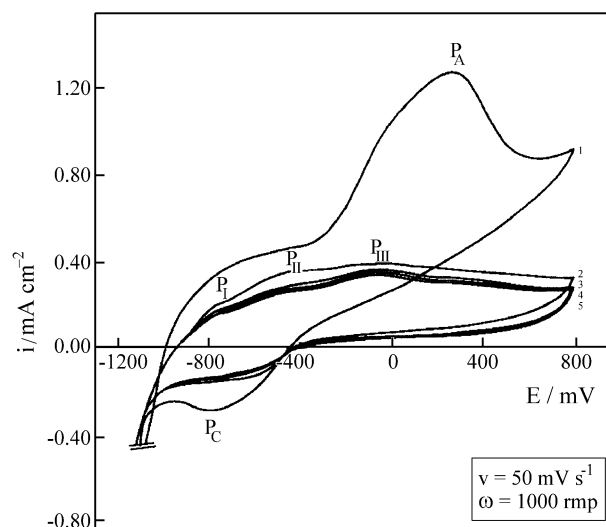


Figure 3. Cyclic voltammograms of iron in 0.1 mol/L sodium perchlorate + 2500 ppm borax (pH 9). P_A and P_C : anodic and cathodic peaks in the first cycle. P_I , P_{II} and P_{III} : anodic shoulders in following cycles.

solution containing 2500 ppm of borate. The results shown in Figs. 2 and 3 allow the demonstration of the inhibiting properties of borate ion. As its concentration increases, passivation becomes more efficient with a smaller anodic charge transport. The comparison between the first cycle and the following ones shows that the anodic peak P_A is substituted by three anodic shoulders, P_I , P_{II} and P_{III} (Fig. 3), with smaller current densities. The second cycle and the following ones are identical in shape to the first cycle obtained for iron in 0.1 mol/L borax solution (Fig. 4) under the same conditions, which are associated with passive film formation. Thus, it seems to exist competitiveness among the actions of the different anions: if borate cannot avoid metallic dissolution and the consequent film formation from species precipitation in perchlorate solution, at least it can make it more difficult, polarizing the process that occurs at P_A . In the following cycles, the formed film is not totally reduced, and suffers an ageing processes through dehydration, which makes it more resistive. This film acts as a physical barrier regarding the dissolution of iron in the presence of perchlorate.

The experiments using RRDE, with the ring electrode either kept under oxidizing (+600 mV) or under reducing conditions (-800 mV) during the disc potential scan, allowed the verification of the influence of borate on the dissolution processes. Figure 5 shows the results obtained in solutions of 1000, 2000 and 3000 ppm of borate for the first cycle of the disc. In the forward scan, Fe^{2+} species result from the dissolution of the disc at -650 mV, that corresponds to the active region. The amount of Fe^{2+} species tends to decrease as borate concentration increases. Also Fe^{3+} species are formed at -630 mV. No corresponding current has been detected by the ring in the solution containing the highest concentration of borate (3000 ppm). In the reverse scan, the ring registers Fe^{2+} dissolution during the film reduction, at a more negative potential than in the

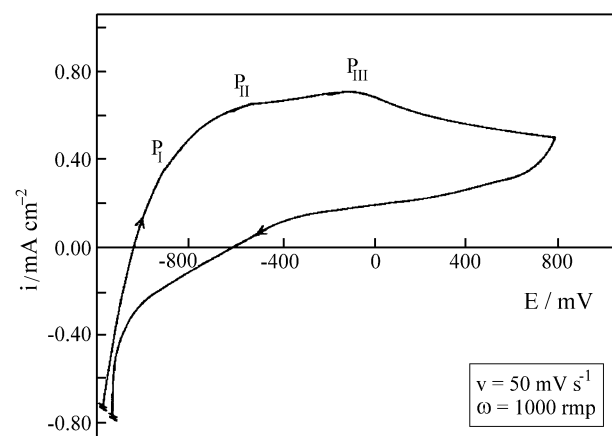


Figure 4. Cyclic voltammogram of iron in 0.1 mol/L borax solution (pH 9). P_I , P_{II} and P_{III} : anodic shoulders. P_C : cathodic current region.

absence of borate. The ring electrode current also diminishes as borate concentration increases, as in the direct scan.

The results obtained with the ring-disc electrode permit the comparison of the metal dissolution processes in the absence as well as in presence of borate. This anion seems to prevent the dissolution of iron as Fe^{2+} and Fe^{3+} species at potentials higher than -400 mV, exactly when there is a greater dissolution in the medium containing only perchlorate. Thus borate stabilizes a film formed at more negative potentials, corresponding to the active zone, and also helps the film in preventing the dissolution process at higher potentials (Fig. 1.b, peak 3).

The charge, evaluated by integrating the ring current along the oxidation process, permits the calculation of the fraction of the total charge transferred at the disc referring to the dissolution processes. Thus, in the solution containing 1000 ppm borate, this fraction corresponds to 1.7% and diminishes to only 1.1% in solutions containing 2000 or 3000 ppm. This fact shows that the borate inhibiting action on the dissolution processes does not increase remarkably with higher concentrations.

The influence of pH on the electrochemical behavior of iron in perchlorate+borate solutions has also been verified. The borate concentration was fixed at 1000 ppm, potential scan rate was kept at 50 $mV s^{-1}$. pH values from 9 to 12 have been tested by adding NaOH to the solutions. As the pH increases, the current referring to peaks P_A and P_C decreases. The anodic peak potential is shifted to more positive values and the cathodic peak to more negative values. The increase of pH causes a polarization of the processes represented by peaks P_A and P_C , increasing thus the redox process irreversibility. The anodic and cathodic

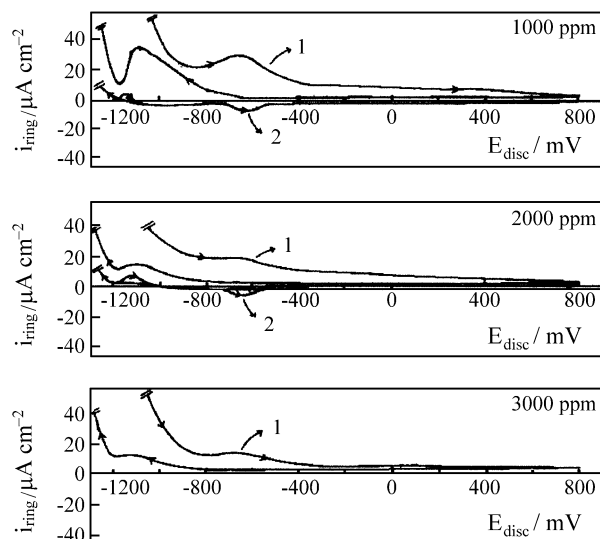


Figure 5. Platinum ring current density variation with disc potential scan in 0.1 mol/L sodium perchlorate with borax addition (pH 9). 1 - ERING = + 600 mV; 2 - ERING = - 800 mV.

current densities also decrease as pH increases. The presence of oxydriles improves borate inhibiting action.

Figure 6 presents the effects of borate concentration and pH on the density of the anodic current peak (i_{PA}) and of the corresponding peak potential (E_{PA}). The variation of $\log i$ with $\log C_{BORATE}$ is also shown, where i is the anodic current density at -100 mV in the Tafel region of the anodic peak P_A (Fig. 3). The straight line obtained has a linear correlation coefficient of 0.96 and an inclination of -0.31, which corresponds to the reaction order of the oxidation process with respect to the borate anion. The negative value indicates the inhibiting character of this anion. From current densities evaluated at the same potential (-100 mV) but for different pH with fixed borate (1000 ppm) and perchlorate

(0.1 mol/L) concentrations, a value of -0.20 was found for the reaction order in relation to the pH, showing that this factor also contributes to the inhibition of the metal oxidation, but in a less remarkable way.

The surface fraction covered with borate species has been evaluated for iron in perchlorate solutions containing different amounts of borate and data were fitted to a Langmuir isotherm. This isotherm is defined by the expression:

$$\frac{C_{borate}}{\theta} = \frac{1}{K_{adsorption}} + C_{borate} \quad (1)$$

where C_{BORATE} is the borate anion concentration in mol/L; $K_{adsorption}$ is the equilibrium constant of the adsorption / desorption process and θ is the fraction of the electrode surface covered by the adsorbed species. This θ can be determined by:

$$\theta = \frac{(i_o - i_i)}{i_o} \quad (2)$$

where i_o is the anodic current density in absence of borate and i_i is the anodic current density in different concentrations of this anion, evaluated at a constant potential (-200 mV) before the anodic peak P_A . From Figs. 1 and 2, i_o and i_i for different borate concentrations have been determined and the values for θ as a function of C_{BORATE} were calculated. Figure 7 represents C_{BORATE} / θ

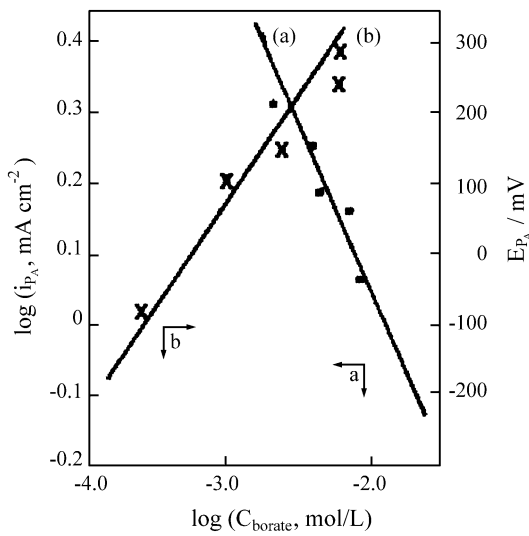


Figure 6.1. a. Variation of anodic peak current density with borax concentration. b. Effect of borax concentration on the anodic peak potential.

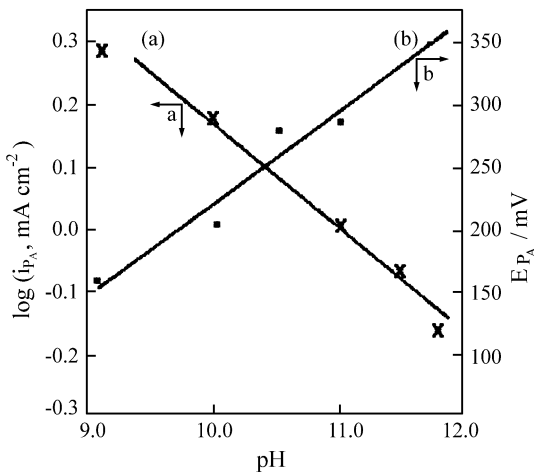


Figure 6.2. a. Anodic peak current density as a function of pH. b. Variation of anodic peak potential with pH.

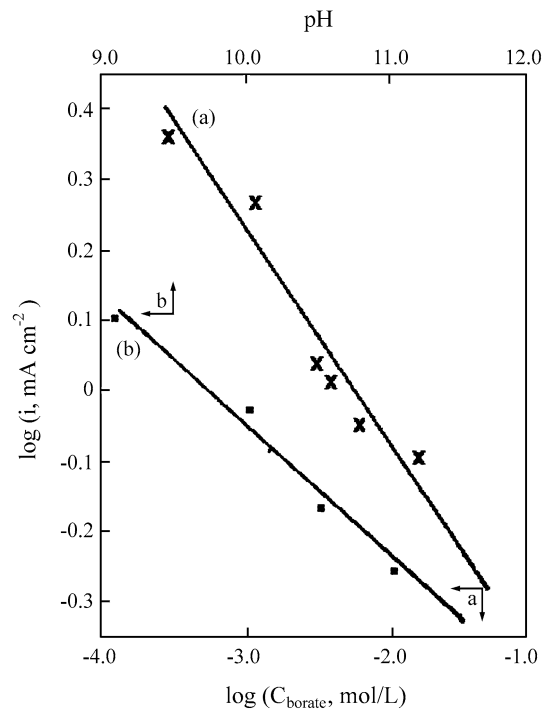


Figure 6.3. a. Variation of anodic current density (at $E = -100$ mV) with borax concentration. b. Variation of anodic current density (at $E = -100$ mV) with pH.

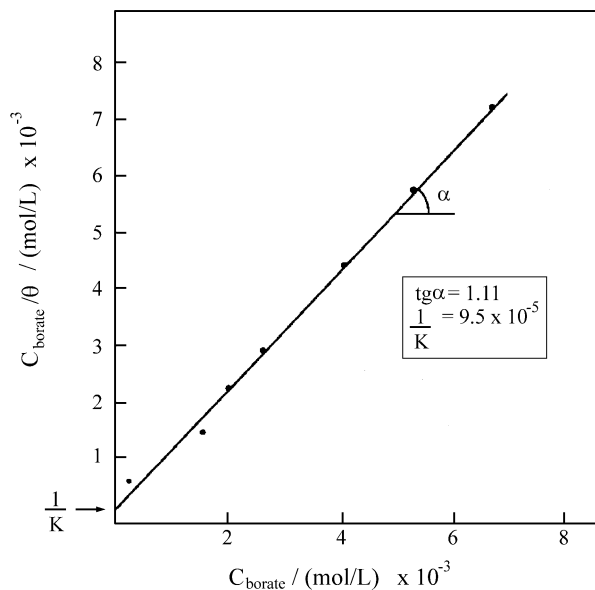


Figure 7. Langmuir isotherm for iron in 0.1 mol/L sodium perchlorate (pH 9) with different borax concentrations.

against C_{BORATE} . A linear behaviour is found as is foreseen by the Langmuir isotherm with a linear correlation coefficient of 0.9999. The straight line slope value (1.11) is near to the theoretical value (1.0). The linear coefficient ($1/K_{\text{adsorption}}$) permits the evaluation of $K_{\text{adsorption}}$, resulting in a value of 1.05×10^4 . The free energy variation for the adsorption process defined as $\Delta G^0 = -RT \ln K_{\text{adsorption}}$, resulted in -5.5 kcal/mol for the experimental conditions of this paper. This value indicates that the interaction between borate ion and the electrodic surface occurs by chemical adsorption.

Conclusions

The electrochemical behaviour of iron in the slightly alkaline solutions investigated depends on the anion present in solution. In the presence of perchlorate, one anodic peak (P_A) and one cathodic peak (P_C) are observed. In the presence of only borate, three anodic shoulders (P_I , P_{II} and P_{III}) and one cathodic current region are observed with smaller current densities than in the perchlorate medium. These anodic shoulders in the medium containing only borate may be related to film formation reactions with different kinetic controls. Small amounts of borate added to perchlorate solution change the voltammetric profile. The anodic peak P_A as well as the cathodic peak P_C diminish and their potentials are shifted to more positive values. Thus there is a meaningful decrease of the total anodic charge for passivation of iron and an increase of the passivation efficiency in this medium. During the potentiodynamic experiments, when the electrode potential is cycled, the remaining film which had not been totally reduced in the previous

scan, makes the passivation even more efficient, probably as it offers nucleation sites for the continuous growth of the oxide film. The RRDE experiments showed that borate anions, even in small quantities, inhibit iron dissolution. This property is due not only to its buffering capacity, that maintains the alkaline pH close to the electrode, but also to its chemical adsorption on the metallic surface, according to a Langmuir isotherm, stimulating the formation of a better film of passive oxide that obstructs dissolution.

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Received: November 9, 1998