

Selective Sorbent Enrichment of Nickel Ions from Aqueous Solutions using a Hierarchically Hybrid Organic-Inorganic Polymer Based on Double Imprinting Concept

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Este estudo envolve a preparação de um polímero orgânico-inorgânico impresso com íon (IIP) para o enriquecimento seletivo de Ni²⁺ em soluções aquosas com posterior determinação por GF AAS. A rede polimérica híbrida foi preparada a partir de uma mistura do monômero funcional, 2-aminoetil-3-aminobutilmetildimetoxissilano e tetraetoxissilano como agente de ligação cruzada, contendo Ni²⁺ e brometo de cetiltrimetilamônio como moldes. Os polímeros foram caracterizados por IV, TG e MEV. A capacidade máxima adsorptiva do IIP para os íons Ni²⁺, determinada pelo modelo linear de Langmuir, foi de 5,44 mg g⁻¹. Quando o coeficiente de seletividade iônica do IIP foi comparado com os coeficientes de seletividade do NIP (non imprinted polymer) e IIP2 (ion imprinted polymer na ausência do surfactante), a partir das misturas binárias de Ni²⁺/Cu²⁺, Ni²⁺/Co²⁺, Ni²⁺/Cd²⁺ e Ni²⁺/Zn²⁺, os respectivos valores do coeficiente de seletividade relativo (k') foram 36,54 e 3,55, 1,22 e 2,03, 4,43 e 1,42, 28,60 e 1,74, demonstrando maior seletividade do IIP para os íons Ni²⁺. O método proposto forneceu um limite de detecção de 0,16 µg L⁻¹ e foi aplicado com sucesso na determinação de Ni²⁺ em amostras de águas enriquecidas e fitoterápico (*Gingko Biloba*) com valores satisfatórios de recuperação.

This study involves the preparation of an ion imprinted organic-inorganic polymer (IIP) for selective sorbent enrichment of Ni²⁺ from aqueous solutions with further determination by GF AAS. The hybrid polymeric network was prepared from a mixture of 2-aminoethyl-3-aminobutylmethyldimethoxysilane as functional monomer, and tetraethoxysilane as crosslinking agent, containing Ni²⁺ and cetyltrimethylammonium bromide as templates. The polymers were characterized by IR, TG and SEM. The maximum adsorptive capacity of IIP towards Ni²⁺ ions, determined by linear Langmuir model, was found to be 5.44 mg g⁻¹. When the selectivity coefficient of IIP was compared with the selectivity coefficient of NIP (non imprinted polymer) and IIP2 (ion imprinted polymer in the absence of surfactant), from the binary mixtures of Ni²⁺/Cu²⁺, Ni²⁺/Co²⁺, Ni²⁺/Cd²⁺ and Ni²⁺/Zn²⁺, values of relative selectivity coefficient (k') were 36.54 and 3.55, 1.22 and 2.03, 4.43 and 1.42, 28.60 and 1.74, respectively, demonstrating higher selectivity of IIP for Ni²⁺ ions. The proposed method provided a limit of detection of 0.16 µg L⁻¹ and was successfully applied for Ni²⁺ determination in spiked water samples and in a phytoterapeutic product (*Gingko Biloba*) with satisfactory recovery values.

Keywords: nickel, Hierarchically Imprinted Polymer, GF AAS, enrichment

Introduction

Recent years have seen remarkable progress in the development of separation/preconcentration methodologies

for analysis of sample at low analyte levels, mainly for solving instrumental problems. Indeed, solid phase extraction (SPE) has been the most used technique for this purpose, due to its natural features such as high reproducibility, high preconcentration factor, ease of automation and easy regeneration of solid phase.¹ A survey

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on the literature shows that the primary requirements to attain good performance of the SPE include quantitative sorption and elution, kinetically faster sorption and desorption, accessibility of analyte towards sites of sorbents and, mainly, selectivity.² The selective analyte separation has not been properly achieved from analysis of real samples by using conventional solid sorbents, such as modified silica (C₁₈), chelating resins, unloaded polyurethane foam and activated carbon, thus being the main disadvantage of the SPE.³ Nevertheless, this drawback can successfully be overcome by using selective sorbents, which present specific recognition sites formed in synthetic materials by binding analyte with high specific sites of sorbent.⁴ These materials make use of chemical imprinting technology that provides selective recognition at molecular level of an analyte over closely related structural analogues. Molecularly imprinted polymers have been the most common approach of chemical imprinting technology and have found application in several analytical methods such as, chromatographic separation, sensors, solid phase extraction for sample clean-up and fluorescence spectroscopy.⁵ Based on simple specific memory effect created as a consequence of the shape and the position of functional group of the template molecules with those functional groups of sorbent, Nishide *et al.*,⁴ in a similar way, proposed, for the first time, sorbent materials capable of recognizing metallic ions. The selectivity of these ion imprinted materials depends on the specificity of the ligand, coordination geometry, coordination number charges and sizes of the ions.^{6,7} Thus, it is of paramount importance the preparation of these materials for the development of selective sorbent enrichment methods for metallic ions. In fact, similarly to molecularly imprinting materials, ion imprinted materials find interesting applications in SPE, metal ion sensors and membranes separation of inorganics, but their use is significantly less exploited. Different approaches in synthesis of these sorbents can be adopted and mainly categorized as organic or hybrid organic-inorganic sorbents. In organic polymeric sorbents, the “memorized cavity” by means binding metal ion templates with those sites of organic monomers, can be achieved from different methods, such as crosslinking of bifunctional reagents with linear polymeric chains, chemical immobilization, surface imprinting and trapping. In order to get more information, the reader can refer to a review described by Rao *et al.*² Regarding organic-inorganic polymeric sorbents, the conventional sol-gel process has widely been employed for the incorporation of the metal ion templates into rigid networks⁸, using tetraethoxysilane (TEOS) as the inorganic precursor and organic molecules as functional monomers. In addition, ion imprinted sorbents can also be prepared

by surface imprinting technique, in which the silica gel surface is modified with organic molecules as functional monomers.⁹ Although these two conventional methods provide sorbents with attractive features, including high selectivity and absence of swelling effect,^{9,10} both do not present a precise control of pore sizes and surface area.¹¹ Moreover, especially for sol-gel process, the resulting materials exhibit reduced accessibility of the analyte towards linkage sites, promoting unfavorable adsorption/desorption kinetics. Therefore, improved hybrid organic-inorganic sorbents with a precise control of not only adsorption sites but also with mesoporous structures have been developed by using surfactants during the synthesis.¹² These sorbents are known as hierarchically imprinted organic-inorganic sorbents because they utilize the imprinting effect twice (metal ion template and surfactant). Their application as selective materials are still in early stages and have been limited only for copper,¹² cadmium and mercury,^{13,14} whose functional employed groups have been (3-(2-aminoethylamino)-propyltrimethoxysilane) and *N*-[3-(trimethoxy-silyl)propyl]ethylenediamine (TPED).

Thus, in the present study, a hybrid organic-inorganic sorbent based on double imprinting concept was evaluated for selective retention of nickel ions from aqueous solution using 2-aminoethyl-3-aminobutylmethyldimethoxysilane (AAMDMS), as functional monomer. The sorbent was characterized by infrared spectroscopy, thermogravimetry (TG), scanning electron microscopy (SEM) and adsorption isotherms. In addition, it was compared with control polymers taking into account the relative selectivity coefficient (k'). Finally, by using the selective sorbent a reliable preconcentration method for nickel determination by GF AAS was developed and its feasibility was assessed by its application for analysis of water samples. The whole optimization procedure was performed by using factorial design and Doehlert matrix owing to its intrinsic advantages over univariate method.¹⁵

Experimental

Instrumentation

A graphite furnace atomic absorption spectrometer (GF AAS) (Zeiss AA55, Jena, Germany), equipped with deuterium lamp for background correction and a hollow cathode lamp as radiation source for nickel, was used. The hollow cathode lamp was operated at 5.0 mA and the wavelength was set at 232.0 nm. Aliquots of 20 μ L of each sample and calibration solution using MPE 5 autosampler were directly injected into the pyrolytic graphite tube without L'vov platform with transversal heating. Argon

99.99% (White Martins, São Paulo, Brazil) was used as purge gas. A Shimadzu AA-6800 flame atomic absorption spectrometer (FAAS) (Shimadzu, Tokyo, Japan) was also used in this study. The pH values of samples were measured using a pH-meter (Scott, Cambridge UK), equipped with a glass-combination electrode. A peristaltic pump (Ismatec IPC-08, Glattzbrugg, Switzerland), furnished with Tygon® tubes, was used to propel all samples and reagent solutions. The morphology of all polymers was evaluated by scanning electron microscopy (SEM), using a JEOL JSM-6360 LV scanning electron microscope (Tokyo, Japan) with an electron acceleration voltage of 20 kV. The samples were previously coated by a thin layer of carbon followed by the Au/Pd alloy in a Bal-Tec MED 020 equipment. Thermal analyses were performed in a TA Instruments TGA 2950 thermogravimetric analyzer (California, USA), where the polymers (*ca.* 10 mg) were heated from 30 up to 980 °C, at a heating rate of 10 °C min⁻¹, under argon atmosphere. Infrared (IR) spectra were obtained in a Bomem Michelson MB-102 spectrometer. The experiments were carried out using KBr discs to prepare the polymer samples and the spectral range varied from 4000 to 400 cm⁻¹. Sample decomposition was performed in a microwave oven (Milestone Microwave Laboratory System).

Reagents and solutions

The solutions were prepared with analytical grade chemical reagents as well as with water obtained from a Milli-Q purification system (Millipore, Bedford, MA, USA). In order to prevent metal contamination from laboratory glassware, it was kept 24 h in a 10% v/v HNO₃ solution. Tetraethoxysilane (TEOS) and 2-aminoethyl-3-aminobutylmethyldimethoxysilane (AAMDMS) from Alfa Products were used in this study. Cetyltrimethylammonium bromide (CTAB) (Vetec, Brazil) was used without further purification. Nickel standard solutions were prepared daily by appropriate dilution of 1000 mg L⁻¹ nickel solution (Merck, Darmstadt, Germany). Acetate, borate, ammoniacal and phosphate buffers were prepared from their respective salts (Merck) without further purification. Ethanol of analytical grade was purchased from Merck. All solutions used in interference study containing the Cu²⁺, Cd²⁺, Co²⁺, Sb³⁺, Pb²⁺, Mn²⁺, Zn²⁺ foreign ions were prepared by dilution from Merck standard solutions at 1000 mg L⁻¹.

Preparation of hierarchically hybrid organic-inorganic polymer

The synthetic route and the recognition mechanism of Ni²⁺ ions were based on literature with minor modification.¹³

In the first step, the complex formation was performed between Ni²⁺ ions and those binding sites (-NH- and -NH₂) present in AAMDMS in the presence of CTAB. For this task, 0.550 g of NiCl₂·5H₂O and 1.46 g of CTAB were dissolved in 40 mL of deionized water. To this mixture, 1.1 g of AAMDMS was added and magnetically stirred for 1 h. Next, TEOS (4.0 g of TEOS was dissolved in 4 mL of methanol and stirred for 20 min) and NaOH (1.0 mol L⁻¹, 7.6 mL) were added to the mixture, which was stirred for 2 days at room temperature and, then, refluxed at 90 °C for 1 day. In this procedure, TEOS and AAMDMS were independently self-hydrolyzed and self-condensed to form the hybrid organic-inorganic polymer. The hydrolysis of TEOS and the copolymerization occur due to the presence of NaOH. After the end of the synthesis, Ni²⁺ ion confined in the selective cavity and the surfactant incorporated in the sorbent were removed from successive washings with 2.0 mol L⁻¹ HNO₃ and ethanol, respectively. The resulting polymer was exhaustively washed with deionized water, and dried at 100 °C for 4 h until use. In order to evaluate the selective features of the polymer as well as the effect of CTAB in the synthesis, the Ni²⁺-imprinted polymer, defined here as IIP2, was also prepared using the same synthesis procedure, except by addition of CTAB. In addition, the control blank polymer, known as NIP (non imprinted polymer) was prepared in parallel without addition of Ni²⁺, however containing CTAB. These polymers (IIP, NIP and IIP2) were further characterized by SEM, thermogravimetry (TG) and IR.

General procedure for enrichment of nickel ions from water samples

The ion imprinted polymer (IIP) (100 mg) was packed into a minicolumn (6.0 x 1.0 cm i.d.), made of poly(vinyl chloride) (PVC). Small amounts of glass wool were placed at both sides in order to prevent loss of the polymer during sample loading. Afterwards, the minicolumn was successively conditioned with 3.0 mol L⁻¹ HNO₃ and deionized water. The method was based on off-line enrichment of 20.0 mL of sample previously buffered (pH 8.4) with 0.43 mol L⁻¹ borate buffer onto polymer at a flow rate of 5.0 mL min⁻¹. After preconcentration procedure, the elution step was carried out with 2.0 mL of 3.3 mol L⁻¹ HNO₃ solution at a flow rate of 2.0 mL min⁻¹. The eluate was collected into an autosampler cup and an aliquot of 20.0 µL was introduced into graphite furnace without adding chemical modifier. All operation conditions of GF AAS were employed according to manufacturer's instructions. The measurements were evaluated by peak area of the absorbance signals.

Batch adsorption experiments

Adsorption experiments were carried out in closed polyethylene flasks at room temperature, by putting 30 mg of IIP in 20.0 mL of Ni²⁺ ions solutions with increasing concentration ranging from 3.0 up to 11.0 mg L⁻¹. The value of optimum pH (pH 8.4 in 0.43 mol L⁻¹ borate buffer) adopted in these experiments was that previously optimized from minicolumn preconcentration procedure, while shaking time employed was 10 min. After the shaking time, the concentration of Ni²⁺ in the supernatant was determined by FAAS. The amount of Ni²⁺ adsorption on the IIP was determined as the difference between the initial concentration and the equilibrium concentration (supernatant). In order to achieve the maximum adsorption capacity (MAC), all batch experimental data were fitted to the linear Langmuir isotherm model.¹⁶

Optimization of enrichment method

The selective sorbent enrichment procedure of nickel ions was optimized by using a 2⁵⁻¹ fractional factorial design associated with Doehlert matrix. These techniques were adopted in this study because they enable the selection of optimal experimental procedures, helping to avoid trivial mistakes during the optimization. The possibilities of simultaneous study of several factors also facilitate the evaluation of interaction effects between the factors as well as allow us to attain the best sensitivity of the method. Those factors commonly studied, which play important role in the sorbent enrichment system were: sample pH, eluent concentration (EC), buffer sample concentration (BC), enrichment flow rate (EFR) and mass of IIP. All assays were performed by enrichment of 20.0 mL of 10.0 µg L⁻¹ nickel solution. The analyses of experimental assays were processed using the STATISTICA software package (StatSoft, Tulsa, USA).

The eluent flow rate was fixed at 2.0 mL min⁻¹. It must be emphasized that higher amounts than 200 mg of polymer were avoided throughout the study, once this condition leads to leakage in the minicolumn, due to resulting overpressure.

Evaluation of imprinting effect on the IIP

In order to measure the specificity of the IIP towards Ni²⁺ ions, adsorption studies under competitive conditions in the presence of Cu²⁺, Co²⁺, Cd²⁺ and Zn²⁺ ions were performed. As it is predictable that in the presence of foreign ions a competition might start for those sites of IIP, the imprinting effect of IIP towards Ni²⁺ ions was

investigated from competitive adsorption of binary mixtures of Ni²⁺/Cu²⁺, Ni²⁺/Co²⁺, Ni²⁺/Cd²⁺, Ni²⁺/Zn²⁺ onto IIP, IIP2 and NIP. For this task, batch experiments were carried out at room temperature, where 100 mg of polymers were taken and stirred for 30 min in closed polyethylene flasks with 50 mL of binary mixtures, containing 5 mg L⁻¹ of each metallic ion, buffered (pH 8.4) with 0.43 mol L⁻¹ borate buffer (pH was adopted here according to the optimization procedure). After this step, the supernatant was collected and the metallic ion amounts were determined by FAAS. Based on these assays, parameters related to the selective performance of sorbents, such as distribution coefficient (K_d), selectivity coefficient (k) and relative selectivity coefficient (k') were determined. Distribution coefficients of Ni²⁺, Cu²⁺, Co²⁺, Cd²⁺ and Zn²⁺ were calculated according to equation (1):

$$K_d = \left[\frac{(C_i - C_f)}{C_f} \right] \times \frac{V_s(\text{mL})}{\text{mass of IIP (g)}} \quad (1)$$

where C_i, C_f and V_s represent the initial, final solution concentrations and the volume of the solution, respectively. The selectivity coefficient is defined as the ratio between the distribution coefficient for Ni²⁺ ions and the distribution coefficient for foreign ions. In order to make a comparison between k values of the imprinted polymer with those control polymers, *i.e.* the relative selectivity coefficient (k'), the following equation (2) was employed. This equation allows an estimation of the effect of imprinting on selectivity.

$$k' = \frac{K_{\text{imprinted}}}{K_{\text{control}}} \quad (2)$$

Results and Discussion

Characterization of hierarchically hybrid organic-inorganic polymer

IR spectra of IIP, NIP and IIP2 are illustrated in Figure 1. As it can be seen from the spectra, the vibrations of OH relative to adsorbed water were observed around 3440 and 1627 cm⁻¹. The absorption around 1380 cm⁻¹ is assigned to CH₂-N vibrations and a stronger and broader peak can be observed from IIP and IIP2 spectra, in which this phenomenon may be explained by coordination between -CH₂-N- and Ni²⁺ ions, still retained into the selective cavities of IIP. In fact, imprinted polymer has an inherent limitation in removal of total template from selective cavities. In the three spectra, the observed peaks around 1079 and 947 cm⁻¹ are attributed to the stretching

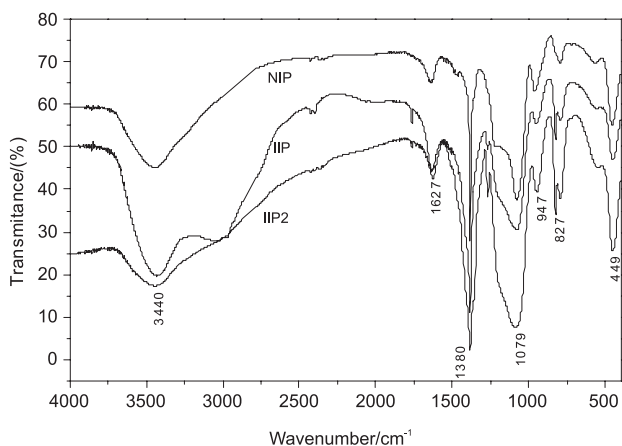


Figure 1. FT-IR spectra of polymers IIP, NIP and IIP2

of Si–O–Si and Si–O–H bonds, respectively, and the peaks around 827 and 449 cm^{-1} correspond to the Si–O vibrations. The presence of two peaks of very low intensity around 3030 and 2969 cm^{-1} from IIP spectrum is attributed to the symmetry and asymmetry stretch vibration of CH_2 of CTAB molecule.¹⁷

The morphology of the polymers is shown in Figures 2a-c. The images clearly shown the presence of pores in the IIP network and significant changes in the morphology of the IIP and NIP samples were observed. The IIP (Figure 2a) was found to have a larger surface area and particles with more uniform sizes than IIP2 (Figure 2c). This behavior demonstrates that the presence of surfactant in the synthesis may promote a better access of Ni^{2+} ions towards the selective cavity. On the other hand, NIP presented completely flat morphology, thus indicating that the presence of Ni^{2+} ions during the synthesis is fundamental and play a more important role than the surfactant for the formation of cavities and pores in the final product.

Thermogravimetric curves for the three polymers are shown in Figure 3. The polymers studied here are classified as polysilsesquioxane and their thermal degradation take place via scission and redistribution of Si–C and Si–O bonds, where high decomposition temperatures are observed, commonly accompanied by formation of silicon oxycarbide glasses.¹⁸ According to Figure 3, the initial weight loss from 25 up to around 180 $^{\circ}\text{C}$ was attributed to the physisorbed residual water and loss of volatile segments retained in the hybrid on account of incomplete polycondensation reactions of the residual silanol groups. For the IIP, a pronounced weight loss from 214 up to 223 $^{\circ}\text{C}$ (*ca.* 49%) was verified due to the thermal decomposition of the polymer and the subsequent weight loss (about 11%) was detected in the range of 224–950 $^{\circ}\text{C}$, in which it was associated to the formation of silicon oxycarbide

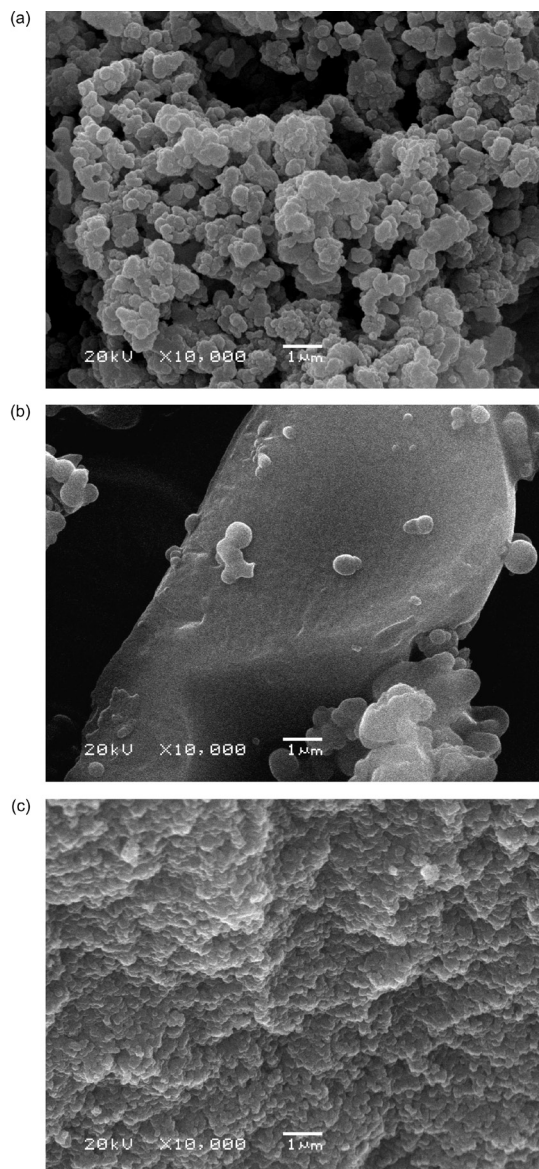


Figure 2. Scanning electron micrographs of (a) IIP, (b) NIP and (c) IIP2. The images were 10,000 amplified.

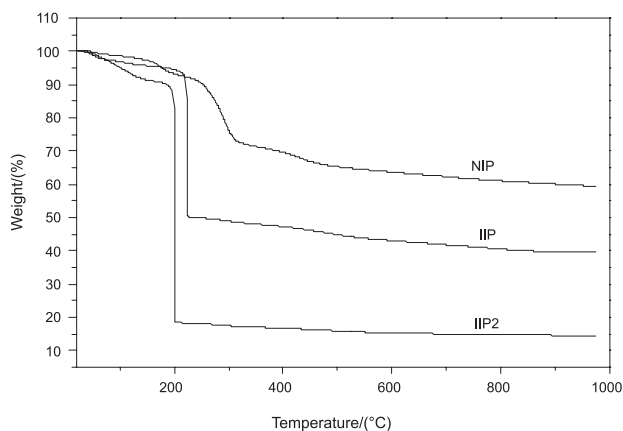


Figure 3. Thermogravimetric curves of polymers IIP, NIP and IIP2.

glasses, which are stable at higher temperatures. A similar behavior was observed for the IIP2; however the polymer degradation was verified from 189 up to 201 °C, with a weight loss of 81%. Hence, the higher thermal stability and ceramic yield of IIP in relation to IIP2 indicated that there is a synergic effect between Ni²⁺ ions and the surfactant in the thermal stability of the IIP, in which it promoted a high crosslinking degree of the polymer. The thermal behavior of NIP showed a different degradation pattern when compared with IIP and IIP2. From 25 up to 248 °C, the weight loss (ca. 9 %) is attributed to physisorbed residual water and loss of volatile segments. Afterwards, the pyrolysis was obtained when the temperature varied from 248 up to 315 °C. In this step, the weight loss occurred by degradation of Si-C bonds and redistribution of Si-O. Finally, the final step carried out up to 950 °C is assigned to the stable silicon oxycarbide glasses. A plausible explanation of the expressive difference of thermal stability and ceramic yield of NIP when compared with IIP and IIP2 is related to the porosity of the material. As already demonstrated,

micrographs showed that NIP network do not exhibit porosity. Therefore, it is proposed that to the formation of densely packed crosslinked networks favors the increase of the initial temperature of polymer decomposition.

Optimization of enrichment procedure

Table 1 shows the factors, their levels and the analytical response for the 2⁵⁻¹ fractional factorial design. For this design, the generator matrix Mass = pH.EC.BC.EFR was adopted. The main effects and their interactions can be graphically represented from the Pareto chart (Figure 4). Based on main effect, the illustration shows that only the sample pH is statistically significant. Both buffer concentration and the eluent concentration are significant factors when analyzed from their interaction effects. On the other hand, the mass of IIP and the enrichment flow rate did not show to be significant for the proposed method. A mass of IIP of 100 mg was, therefore, employed in this study for all other optimization steps. Even working with

Table 1. 2⁵⁻¹ fractional factorial design and responses for the nickel enrichment on IIP

	Factors				Levels		
					Minimum (-)		Maximum (+)
	pH				5.0		8.0
	Eluent concentration (EC) (mol L ⁻¹)				2.0		4.0
	Buffer concentration* (BC) (mol L ⁻¹)				0.05		0.1
	Enrichment flow rate (EFR) (mL min ⁻¹)				5.0		10.0
	Mass of IIP (mg)				100		200
Runs	pH	EC	BC	EFR	Mass	Response (peak area)	
1	-	-	-	-	+	0.252/0.261	
2	+	-	-	-	-	0.593/0.591	
3	-	+	-	-	-	0.125/0.121	
4	+	+	-	-	+	0.517/0.572	
5	-	-	+	-	-	0.081/0.083	
6	+	-	+	-	+	0.864/0.884	
7	-	+	+	-	+	0.153/0.140	
8	+	+	+	-	-	0.718/0.682	
9	-	-	-	+	-	0.122/0.113	
10	+	-	-	+	+	0.812/0.847	
11	-	+	-	+	+	0.210/0.202	
12	+	+	-	+	-	0.723/0.698	
13	-	-	+	+	+	0.169/0.165	
14	+	-	+	+	-	0.749/0.774	
15	-	+	+	+	-	0.047/0.044	
16	+	+	+	+	+	0.924/0.941	

* Acetate/acetic acid buffer for pH 5.0 and phosphate for pH 8.0

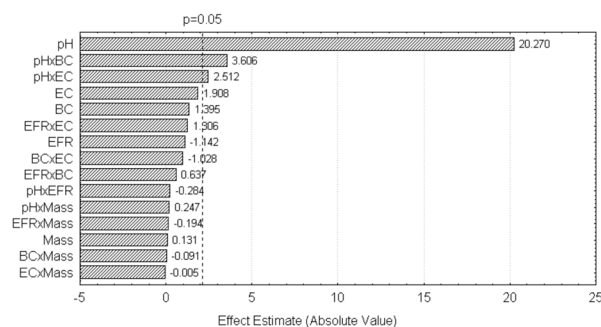


Figure 4. Pareto chart used for assessing the effect of factors on pre-concentration procedure. EC = eluent concentration; BC= buffer concentration; EFR = enrichment flow rate; Mass = mass of IIP.

high enrichment flow rate (10 mL min^{-1}), the analytical signal was satisfactory, indicating that the kinetic of mass transfer of Ni^{2+} towards selective sites of IIP is fast and independent of flow rate. Therefore, 10 mL min^{-1} was chosen for subsequent studies.

An enhancement of Ni^{2+} adsorption at the IIP was noted by increasing the sample pH within the experimental domain (5.0 to 8.0). At lower pH values, the nitrogen atoms of the functional precursor are positively charged and, as a consequence, an electrostatic repulsion occurs with Ni^{2+} ions. In addition, it is interesting to note that the analytical response is increased when the sample pH and the buffer concentration are simultaneously increased. This observation can be evidenced by significant positive interaction effect (3.60). So, a higher analytical signal will be achieved by simultaneous increase in the sample pH and the buffer concentration levels and such conditions could never be achieved if a univariate optimization was used. In order to evaluate the effect of the eluent concentration, we can take into account the interaction between pH and EC, which is positive and statistically significant (2.51). This result suggests the necessity of using acid concentrations higher than 2.0 mol L^{-1} to enhance the analytical signal as well as to avoid memory effect during the enrichment/elution steps.

In order to find the optimum values for pH, EC and BC, a Doehlert matrix was used. Each factor has now a new minimum, maximum and central level that was chosen bearing in mind the results from 2^{5-1} fractional factorial design. The results obtained are shown in Table 2. From the analysis of these results by using analysis of variance (ANOVA) (Table 3), it can be concluded that the quadratic regression model is significant, once there is no lack of fit with a confidence interval of 95% [$\text{MS}_{\text{lof}}/\text{MS}_{\text{pe}}$ (3.1224) is lower than critical $F_{3,2}$ (19.16)]. An observation of the response surfaces in Figure 5 allows us to obtain optimum values of the studied factors. Hence, based on these results, the following experimental conditions were optimized:

Table 2. Doehlert matrix employed for final optimization of pH, eluent concentration and buffer concentration

Runs	pH	EC (mol L^{-1})	BC* (mol L^{-1})	Response (peak area)
1	8.0	3.0	0.275	0.898
2	8.0	3.0	0.275	0.809
3	8.0	3.0	0.275	0.760
4	8.0	3.0	0.500	0.949
5	10.0	3.0	0.387	0.395
6	8.6	4.0	0.387	0.817
7	8.0	3.0	0.050	0.369
8	6.0	3.0	0.162	0.092
9	7.3	2.0	0.162	0.430
10	6.0	3.0	0.387	0.044
11	7.3	2.0	0.387	0.541
12	10.0	3.0	0.162	0.296
13	9.3	2.0	0.275	0.298
14	8.6	4.0	0.162	0.563
15	6.6	4.0	0.275	0.298

* Acetate/acetic acid buffer for pH 6.0, phosphate for pH ranging from 6.6 up to 7.3, borate for pH 8.6 and ammoniacal pH ranging from 9.3 up to 10.0

Table 3. Analysis of variance obtained from the Doehlert design shown in Table 2

Effects	SS	df	MS	F	p
BC	0.152447	1	0.152447	30.8300	0.030939
BC ²	0.032070	1	0.032070	6.4857	0.125752
pH	0.098972	1	0.098972	20.0156	0.046504
pH ²	0.697517	1	0.697517	141.0618	0.007015
EC	0.024471	1	0.024471	4.9488	0.156093
EC ²	0.100422	1	0.100422	20.3088	0.045878
BC.pH	0.006072	1	0.006072	1.2279	0.383226
BC.EC	0.001931	1	0.001931	0.3906	0.595790
pH.EC	0.027032	1	0.027032	5.4668	0.144344
Lack of fit (lof)	0.046318	3	0.015439	3.1224	0.251945
Pure error (pe)	0.009890	2	0.004945		
Total SS	1.159225	14			

$R^2 = 0.95151$ and R^2 adjusted = 0.86424

8.4 for pH, 0.43 mol L^{-1} for borate buffer concentration and 3.3 mol L^{-1} for eluent concentration.

Batch adsorption experiments

It is well established that to build adsorption isotherms, two important parameters must be optimized: sample pH and equilibrium time. Thus, as previously optimized, the

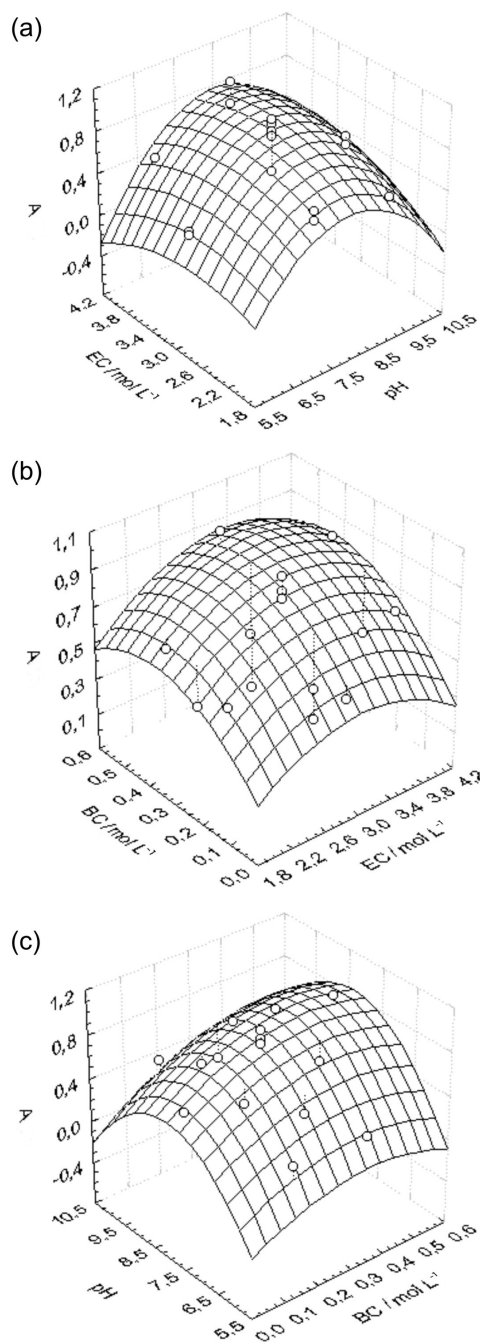


Figure 5. Surface response obtained from Doehlert matrix (a) EC x pH (b) BC x EC (c) pH x BC.

optimum sample pH value of Ni^{2+} adsorption onto IIP was 8.4, while 10 min of shaking time were sufficient for attaining the equilibrium time of Ni^{2+} between the solid and liquid phases (data not shown). By using these values, the adsorption isotherm of Ni^{2+} is illustrated in Figure 6(a). The experimental isotherm data were successfully fitted to the linear Langmuir isotherm model (Figure 6b), with coefficient of determination of 0.9962. The maximum adsorbed amount of Ni^{2+} on the IIP was 5.44 mg g^{-1} . This

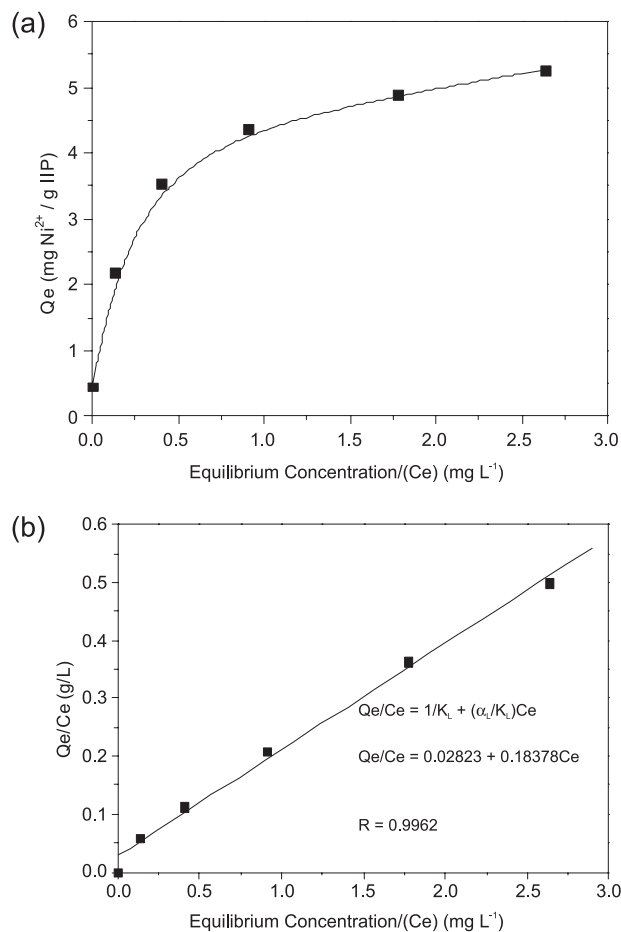


Figure 6. Adsorption isotherm of Ni^{2+} ions on IIP (a), Linear representation of Langmuir equation of Ni^{2+} with IIP.

result shows that the adsorption capacity of the hybrid organic-inorganic polymer towards Ni^{2+} ions is higher than others organic ion imprinted polymers previously published.¹⁹⁻²¹

Evaluation of imprinting effect on the IIP

Tables 4 and 5, show K_d , k and k' values of Cu^{2+} (ionic radius 87.0 pm), Co^{2+} (ionic radius 83.8 pm), Cd^{2+} (ionic radius 109.0 pm) and Zn^{2+} (ionic radius 88.0 pm) with respect to Ni^{2+} (ionic radius 83.0 pm). A brief comparison of K_d values for the Ni^{2+} ions clearly demonstrates high adsorption in IIP and IIP2 in relation to NIP. Such results can be explained due to more porous surface of the IIP and IIP2 in detriment of NIP (Figure 2), which plays an important role in adsorption processes. Similar behavior can be observed for Cu^{2+} , Co^{2+} , Cd^{2+} and Zn^{2+} ions, where low K_d values were observed in NIP. The selectivity coefficient (k) is of paramount importance to express the imprinting effect created in the imprinted polymers. As the stability of Cu^{2+} -diamine complex is higher than Ni^{2+} -diamine complex,¹³

Table 4. Parameters related to imprinting effect (K_d , k , k') of Cu^{2+} and Co^{2+} with respect to Ni^{2+}

Sorbent	K_d (mL g ⁻¹)		k	k'		K_d (mL g ⁻¹)		k	k'	
	Ni^{2+}	Cu^{2+}		IIP/NIP	IIP/IIP2	Ni^{2+}	Co^{2+}		IIP/NIP	IIP/IIP2
IIP	10938	2720	4.02			60795	20678	2.94		
NIP	118	1055	0.11	36.54	3.55	60	25	2.40	1.22	2.03
IIP2	8074	7138	1.13			17365	11975	1.45		

IIP = ion imprinted polymer synthesized in the presence of Ni^{2+} and surfactant; NIP = non imprinted polymer synthesized in the absence of Ni^{2+} and in the presence of surfactant; IIP2 = ion imprinted polymer synthesized in the presence of Ni^{2+} and in the absence of surfactant. The initial concentration of metallic ions was 5.0 mg L⁻¹.

Table 5. Parameters related to imprinting effect (K_d , k , k') of Cd^{2+} and Zn^{2+} with respect to Ni^{2+}

Sorbent	K_d (mL g ⁻¹)		k	k'		K_d (mL g ⁻¹)		k	k'	
	Ni^{2+}	Cd^{2+}		IIP/NIP	IIP/IIP2	Ni^{2+}	Zn^{2+}		IIP/NIP	IIP/IIP2
IIP	93600	20466	4.57			46700	32820	1.43		
NIP	31	30	1.03	4.43	1.42	96	1864	0.05	28.60	1.74
IIP2	68550	21421	3.20			47093	58139	0.82		

IIP = ion imprinted polymer synthesized in the presence of Ni^{2+} and surfactant; NIP = non imprinted polymer synthesized in the absence of Ni^{2+} and in the presence of surfactant; IIP2 = ion imprinted polymer synthesized in the presence of Ni^{2+} and in the absence of surfactant. The initial concentration of metallic ions was 5.0 mg L⁻¹.

it would be expected that larger k values could be obtained for $\text{Cu}^{2+}/\text{Ni}^{2+}$ system in IIP and IIP2, if no imprinting effect was achieved. Nevertheless, as it can be seen in Table 4, the k value for the $\text{Ni}^{2+}/\text{Cu}^{2+}$ system of IIP (4.02) is higher than IIP2 (1.13) and NIP (0.11). Such results confirm that higher recognition sites exist in IIP than IIP2, while no recognition sites are present in NIP. The imprinting effect, quantitatively determined by relative selectivity coefficient, showed that the IIP for $\text{Ni}^{2+}/\text{Cu}^{2+}$ was 36.54 and 3.55 times more selective than NIP and IIP2, respectively. For the $\text{Ni}^{2+}/\text{Co}^{2+}$ system the imprinting effect observed was somewhat reduced, showing relative selectivity coefficient of 1.22 and 2.03 times greater than NIP and IIP2 (Table 4). The competition between Ni^{2+} and Co^{2+} for those selective sites of IIP can be explained due to very similar ionic radius of these metallic ions. Regarding to the $\text{Ni}^{2+}/\text{Cd}^{2+}$ and $\text{Ni}^{2+}/\text{Zn}^{2+}$ binary mixtures, the high adsorption affinity of recognition sites to the imprinted Ni^{2+} ions was once again confirmed from the great relative selectivity coefficient (Table 5).

As demonstrated, IIP presents good imprinting effect when compared with NIP and IIP2. However, in order to apply the enrichment method in real samples, the selective performance of the IIP towards Ni^{2+} ions in the presence of foreign ions [Cu^{2+} , Cd^{2+} , Co^{2+} , Sb^{3+} , Pb^{2+} , Mn^{2+} and Zn^{2+}] at different proportions, was evaluated. For this task, 20 mL of Ni^{2+} ions solutions at 10 $\mu\text{g L}^{-1}$ were preconcentrated under optimized conditions. The studied analyte:interferent ratios were 1:1, 1:10 and 1:100. The tolerance limit is defined

as the foreign ion concentration causing an relative error smaller than 10% in the Ni^{2+} determination. Interferences may be caused by the competitive adsorption of ions onto IIP surface, however only Pb^{2+} interfered at level 100 fold to that Ni^{2+} , leading to a decrease of analytical signal of 17.3%, probably due to Pb^{2+} retention onto those non selective sites of the IIP (Table 6).

Table 6. Effect of foreign ions on the analytical signal recovery of 10 $\mu\text{g L}^{-1}$ Ni^{2+} submitted to preconcentration method

Foreign ions	Recovery Percentage (%)* of the Analytical Signal		
	Ratio analyte:foreign ion		
	1:1	1:10	1:100
Cu^{2+}	100.0	96.4	104.7
Cd^{2+}	98.5	104.0	100.8
Co^{2+}	103.4	94.9	103.9
Sb^{3+}	104.4	99.5	106.7
Pb^{2+}	97.8	100.6	82.7
Mn^{2+}	95.6	103.9	100.0
Zn^{2+}	98.4	96.3	105.5

* Experiments carried out in triplicate

Analytical characteristics

The characteristic performance data for the IIP in the enrichment method under the optimum conditions were obtained by preconcentrating 20 mL of aqueous standard solution of Ni^{2+} ions. The calibration curve for Ni^{2+} ions

was obtained in the range 0.5 up to 15.0 $\mu\text{g L}^{-1}$ with satisfactory coefficient of correlation ($r > 0.994$). A limit of detection ($\text{LOD} = 3 \text{ Std}/m$, where Std is the standard deviation for ten replicates determination of the blank signal and m is the slope of the calibration curve) of 0.16 $\mu\text{g L}^{-1}$ for Ni^{2+} was obtained, while the limit of quantification ($\text{LOQ} = 10\text{Std}/m$) was found to be 0.55 $\mu\text{g L}^{-1}$.²² Eight replicates determinations of both 2.0 and 13.0 $\mu\text{g L}^{-1}$ gave the respective relative standard deviation of 2.67 e 3.37%, respectively. By preconcentrating 20 mL of sample, the experimental enrichment factor of the proposed IIP sorbent, calculated as the ratio between the slopes of analytical curves with and without the enrichment procedure,²³ was found to be 11 fold. Other important analytical parameters for sorbent enrichment were also calculated, such as consumptive index ($\text{CI} = \text{sample volume}/\text{EF}$) of 1.81 mL and enrichment efficiency (EE), which establishes the sensitivity enhancement using a preconcentration time of 1 min. Thus, the EE was found to be 2.75 min^{-1} . These satisfactory results related to the performance of solid sorbent enrichment reflect the outstanding application of

hybrid organic-inorganic polymer in the development of sorbent preconcentration methodologies. Table 7 shows comparative data obtained by proposed method with other off-line sorbent preconcentration for nickel determination. Substantial improvements in sorbent preconcentration performance with the proposed method, such as low sample consumption, absence of chelating agent on the preconcentration procedure, high sample throughput and, mainly, selectivity are shown. It must be emphasized that improvements in the detectability of the method can be achieved by using higher sample volume.

Real sample application

The proposed enrichment method was applied to the determination of Ni^{2+} in water samples. In order to evaluate the feasibility of the method even in samples with high salt content, a synthetic seawater sample was prepared as follows: 27.9 g L^{-1} NaCl, 1.4 g L^{-1} KCl, 2.8 g L^{-1} MgCl_2 , 0.5 g L^{-1} NaBr, 2.0 g L^{-1} MgSO_4 .³² Lake water samples from Furnas Dam, Minas Gerais - Brazil, were collected

Table 7. Characteristic data obtained by selective sorbent preconcentration method using hierarchically hybrid organic-inorganic polymer and other off-line sorbent preconcentration for nickel determination

Sorbent	Preconcentration time (min)	Sample Consumption (mL)	Preconcentration Factor	Samples	Technique	LOD ($\mu\text{g L}^{-1}$)	Ref.
Polycarboxylic microsphere polymer gel	20, 60	100, 300	50, 150	Mineral water, seawater	GF AAS	0.05	24
Silica gel modified with 3-aminopropyltriethoxysilane	2.5	25	5	Seawater	GF AAS	0.5	25
Chromosorb-107 resin	1.25	25	10	Electrolyte manganese samples	GF AAS	-	26
Nickel imprinted microbeads (poly(EGDMA-MAH/Ni(II) microbeads)	100	100	20	Seawater	FAAS	0.3	27
Nickel imprinted organic polymer	100	150	18.7	Water samples	FAAS	1.6	28
Poly(vinyl chloride) modified with 3-ferrocenyl-3-hydroxydithioacrylic acid	120	60	60	Water samples	GF AAS	0.41	29
2-(2-quinolinilazo)-4-methyl-1,3-dihydroxibenzene (QAMDHB) followed by adsorption onto MCI GEL CHP 20Y	30	300	300	Rice, human hair, water samples, soil samples	GF AAS	0.001	30
Nickel imprinted amino-functionalized silica gel	66.6	200	100	Plant samples and water samples	ICP OES	0.16	31
Nickel imprinted organic polymer	25	250	100	Water samples	GF AAS	0.05	20
Nickel imprinted organic-inorganic polymer	4	20	11	Water samples and phytotherapeutic sample	GF AAS	0.16	This work

LOD = limit of detection; GF AAS = graphite furnace atomic absorption spectrometry; FAAS = flame atomic absorption spectrometry; ICP OES = inductively coupled plasma optical emission spectrometry.

in polypropylene bottles, to which 3 drops of concentrated HNO_3 for each 100 mL of samples were added in order to avoid growth of microorganisms. Afterwards, the samples were filtrated under vacuum using 0.45 μm cellulose acetate membranes and buffered with 0.43 mol L^{-1} borate buffer (pH 8.4). Recovery experiments were carried out using sample solutions to which known amounts of Ni^{2+} ions solution had been added. Each recovery was calculated by comparing the results obtained before and after addition. As shown in Table 8 the data obtained for spiked samples showed good recoveries, thus attesting that the sorbent solid phase enrichment associated to GF AAS determination is a reliable way for enrichment of Ni^{2+} in various natural water samples as well as for sample clean-up. The direct analysis of Ni^{2+} in water samples containing high salt content, such as seawater could not be performed by GF AAS, because it generates high background signals, scattering, and a strong matrix effect that lead to a worsening on sensitivity and even loss of accuracy. Apart from the water samples, the method also was applied for Ni^{2+} determination in a phytotherapeutic product (*Ginkgo Biloba*) submitted to acid decomposition according to the literature.¹ The amount of Ni^{2+} ions determined in *Ginkgo Biloba* was found to be $1.05 \pm 0.08 \mu\text{g g}^{-1}$, while the addition of $3.0 \mu\text{g g}^{-1}$ to sample gave rise to a recovery of 92.4% ($3.74 \pm 0.09 \mu\text{g g}^{-1}$).

Table 8. Recovery percentage for nickel ions determination in different kinds of water samples

Samples	Added amount ($\mu\text{g L}^{-1}$)	Determined amount* ($\mu\text{g L}^{-1}$)	Recovery (%)
Tap water	-	ND	-
	2.0	1.80 ± 0.05	90
	4.0	4.36 ± 0.26	109
Lake Water	-	ND	-
	2.0	1.84 ± 0.04	92
	4.0	4.20 ± 0.17	105
Synthetic Seawater	-	ND	-
	2.0	1.99 ± 0.14	99
	4.0	4.04 ± 0.02	101

ND = not detected. *The results are expressed as mean value \pm S.D based on three replicates ($n = 3$) determination. Confidence interval, 95%.

Conclusion

Porous hybrid organic–inorganic polymer based on double imprinting concept for selective retention of nickel ions using 2-aminoethyl-3-aminobutylmethyldimethoxysilane (AAMDMS) as functional monomer was prepared, characterized against control blank polymers and successfully applied for nickel ions enrichment from

aqueous solution. The chemometric tools based on factorial design and Doehlert matrix was successfully applied in the selection of optimal experimental conditions, helping to avoid trivial mistakes during optimization as well as leading to reduction of experiments. The results obtained in this study, based on relative selectivity coefficient, showed that IIP had imprinting effect for the Ni^{2+} in aqueous solution in comparison to IIP2 and NIP. The most selective performance of IIP is attributed to the double imprinting by Ni^{2+} ions and surfactant CTAB during predetermined orientation of those specific binding sites containing functional groups in selective cavities, which it is easily feasible in sol-gel process. It is worth to emphasize that the use of surfactant leads to the formation of polymers with more porous surface only in the presence of template Ni^{2+} , as can be observed from smoother morphology of NIP. It indicates that there is a synergic effect between Ni^{2+} ions and surfactant in the pores formation.

Finally, it could be shown that with the present study, a more expanded application of hierarchically hybrid organic-inorganic polymer based on double imprinting concept in solid phase preconcentration is described, contributing for new investigations of these imprinted polymers in the field of analytical sciences.

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