

Effect of Temperature on Devulcanization of Waste Sidewall Rubber by Supercritical Ethanol

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The influence of temperature on devulcanization of waste sidewall rubber via supercritical ethanol was investigated. The effect of parameters on devulcanizing process was analyzed by a full factorial experimental design. The devulcanized products at different temperatures were characterized by gel permeation chromatography (GPC), differential scanning calorimetry (DSC), thermogravimetry and derivative thermogravimetry (TG-DTG) and Fourier transform infrared (FTIR) spectrophotometry. In addition, the kinetics analysis for the devulcanizing process via supercritical ethanol was established. It was found that the reaction temperature was the only significant parameter. When the temperature rose above 240 °C, the reaction rate increased dramatically. A special phenomenon was observed that the devulcanizing reaction rate of natural rubber (NR) was faster than that of butadiene rubber (BR) in supercritical ethanol. Moreover, when the reaction temperature reached 270 °C, the use of the devulcanizing reagent did not significantly affect the sol fraction. The analysis of experimental results indicated that supercritical ethanol beyond the critical temperature could strongly promote the devulcanizing reaction, maintaining polymer structure of products as the most intact state.

Keywords: supercritical ethanol, waste sidewall rubber, temperature, kinetics

Introduction

The recycling of waste tire rubber has attracted intensive attention due to the great efforts related to the prevention of environmental pollution and conservation of petroleum resource. However, the tire rubber is difficult to be recycled due to the presence of its stable three-dimensional crosslinks. Currently, devulcanization is widely applied in the field of the waste rubber recycling. Through the devulcanizing process, the three-dimensional crosslinked network is broken and the devulcanized waste tire has the ability to reformulate and recycle into usable products.¹ Various types of devulcanizing routes have been studied so far in order to improve the quality of the recycled products, such as chemical method, mechanical process, microwave technique, ultrasonic method and microbial method.²⁻¹³ However, some processes have weaknesses such as environmental pollution and low efficiency.

Supercritical fluid devulcanizing technology is increasingly regarded as an available method for the devulcanization due to its environmental advantage and high efficiency. Crosslinked polymer could be recycled by employing supercritical alcohol. The crosslinked polyethylene has been selectively decomposed by supercritical alcohol and the recycled product could undergo the crosslinking reaction.¹⁴ Furthermore, supercritical alcohol is a kind of solvent which has high chemical activity and could replace the catalyst in some cases. Such as in the degradation process of the alkali lignin, the use of a catalyst did not significantly affect the yields in both solvent systems with alcohol and water.¹⁵ In the field of tire rubber recovery, Kershaw¹⁶ tried to degrade tire rubber to obtain the low molecular weight oil by supercritical ethanol, but most products maintained the macromolecular structure after reaction, which was very close to the devulcanization process. In our previous work,¹⁷ the sidewall rubber was also devulcanized effectively in supercritical ethanol, and the production exhibited the high sol fraction (SF) and lower crosslinking density (CD) of gel.

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As is well known, temperature plays a very important role in the recycling of the crosslinked polymer by supercritical fluid. Increased reaction temperature could increase the amount of radicals and enhance the mobility of monomer, which results in a higher degradation rate of crosslinked network. In ScCO_2 devulcanizing process, the temperature was evaluated as the most significant factor by mathematical statistical method.^{18,19} Meanwhile, the temperature was also the deciding factor for decrosslinking of polymer in supercritical alcohol.²⁰ Under some conditions, increased temperature not only accelerates the degradation of crosslinked network in dynamics, but also changes the activity of the supercritical fluids, which will further promote the reaction. Su *et al.*²¹ achieved the degradation of polyethylene by using supercritical water, and demonstrated that the temperature has significant impact on both the degradation of crosslinked network and the activity of supercritical fluid. However, to the best of our knowledge, few studies about the influence of temperature on the supercritical ethanol devulcanizing process are reported.

In this work, we investigated the role of temperature on devulcanization of the sidewall rubber in supercritical ethanol. A full factorial experimental design was used to analyze the devulcanizing process. The chemical structures and thermal properties of the products were characterized. The reaction rate constant of devulcanizing process was obtained. In addition, a possible reaction path by the effect of temperature was discussed.

Experimental

Materials

The experimental feedstock was sidewall rubber obtained from a passenger car tire (Dunlop tires, model 215/60R16 95H, China). The rough rubber pieces were extracted in the azeotropic acetone/chloroform (3/7 v/v) mixture for 12 h in a Soxhlet extractor to remove the low molecular weight compounds. The extracted pieces were dried to constant weight in vacuum at 50 °C, and then trimmed into strips ($20 \times 3 \times 3 \text{ mm}^3$) as the experimental samples. The experimental samples will be referred to as vulcanizate later. The devulcanizing reagent, diphenyl disulfide (DPDS), was purchased from Energy Chemical, China. Ethanol, acetone, chloroform and toluene were supplied by Sinopharm Chemical Reagent Beijing Co., Ltd, China.

Characterization

Devulcanized product was analyzed by an EQUINOX55 Fourier transform infrared (FTIR) spectrophotometer

(Germany), and the spectrum of sol was taken in the range of 4000 to 400 cm^{-1} with potassium bromide crystals. Gel permeation chromatography (GPC) was performed at 30 °C with a TDA302 chromatograph (USA) to detect the molar mass and molar mass distributions (polydispersity index (PDI)) of sols. The glass transition temperature of vulcanizate and gel was characterized in the range of -100 to 40 °C, with a heating rate of 10 °C min^{-1} in liquid nitrogen by a Q2000 differential scanning calorimeter (DSC) (USA). Thermogravimetry (TG) and derivative thermogravimetry (DTG) were performed in a nitrogen atmosphere with a SETSYS 16/18 thermogravimetric analyzer (France), at a heating rate of 10 °C min^{-1} in the range of 0 to 650 °C.

Experimental process

Devulcanizing reaction was performed in a stainless steel autoclave (100 mL, design pressure of 20 MPa and design temperature of 350 °C). The reaction process is shown in Figure 1. A piece of stainless steel net, which the vulcanizate was placed on, was installed in the middle of the autoclave, and the ethanol was placed on the bottom of the autoclave. Nitrogen was fed into the autoclave, and then discharged for removing the air. This process was repeated five times before starting the reaction. The autoclave was filled with a certain volume of ethanol at ambient temperature, and then followed by heating to the target temperature. The final pressure was referred to as the target pressure, which was only controlled by the volume of ethanol and the target temperature. The heat was provided by an electric heating furnace. The average heating rate was 5 °C min^{-1} . When the reaction was completed, the autoclave was taken out of the furnace and cooled by a blower to ambient temperature.

The devulcanized product was dried to constant weight (W_d) in vacuum at the ambient temperature. If the devulcanizing reagent were added during the

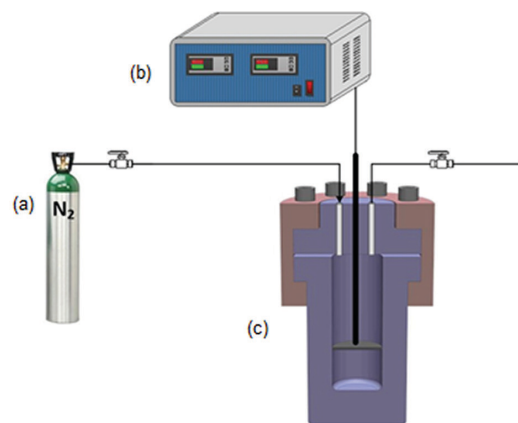


Figure 1. Scheme of the experimental process for devulcanization. (a) Cylinders; (b) control cabinet; and (c) autoclave.

devulcanization process, the dried product would be extracted by acetone for 48 h to remove the unreacted devulcanizing reagent. Then the insoluble fraction was dried again to W_d at the same conditions. Subsequently, the dried product was extracted by chloroform with a Soxhlet extractor for 24 h to obtain the soluble (sol) and insoluble (gel) products. The products were all dried to constant weight in vacuum at the ambient temperature. The sol was comprised of the liner rubber molecules, which were recyclable, and the gel still maintained the crosslinked structure. In general, the higher SF and lower CD of gel reflect a more efficient devulcanized process. The weights of dried sol and gel were W_s and W_g , respectively. The SF was calculated as

$$\text{Sol fraction} = \frac{W_s}{W_d \times w} \times 100\% \quad (1)$$

where w is the weight percentage of rubber ingredient in the sidewall rubber.

The CD of vulcanizate and gel swollen in toluene was calculated by the Flory-Rehner equation:²²

$$v = \frac{-[\ln(1-v_r) + v_r + \mu v_r^2]}{V_0 [v_r^{1/3} - (v_r/2)]} \quad (2)$$

where v is the CD; v_r is the volume fraction of the rubber crosslinked network in the swollen gel; V_0 is the molar volume of the swelling solvent; and μ is the rubber-solvent interaction parameter, $\mu = 0.39$.²³

If the gel contained carbon black, v_r was corrected by equation 3:²⁴

$$\frac{v_{r0}}{v_r} = 0.56^{-z} + 0.44 \quad (3)$$

where v_{r0} is the corrected volume fraction of the rubber

crosslinked network in the swollen gel and z is the mass concentration of carbon black.

Results

Vulcanizate ingredients analysis

The conventional rubber compositions of tires are natural rubber (NR), styrene butadiene rubber (SBR), butadiene rubber (BR) and isobutylene isoprene rubber (IIR).²⁵⁻²⁸ In our previous experiment,¹⁷ we found that IIR was easy to be devulcanized, and even degraded. Therefore, IIR layer was removed before devulcanization experiment.

The ingredient of sidewall rubber was determined by TG-DTG, shown in Figure 2a. A mass loss of about 60% was observed in TG from 300 to 500 °C due to the degradation of the rubber ingredient. The peaks of DTG around 375 and 455 °C represented the maximum mass loss rate of NR and synthetic rubber (SBR or BR), respectively. After fitting analysis, the proportions of NR and synthetic rubber were 53.6 and 46.4%, respectively, in rubber composition. To further analyze the types of synthetic rubber, the sol was extracted from completely devulcanized sidewall rubber as sample for DSC and the analysis result is shown in Figure 2b. It can be observed that two glass transitions occurred at -68 and -101 °C, and the latter corresponded to that of BR. Therefore, the synthetic rubber was BR in the sidewall rubber.

Full factorial experimental design

A full factorial experimental design was used to analyze the influence of the parameters on devulcanizing process. The objective functions were SF and CD of gel. The levels of each factor were chosen by our previous study and are

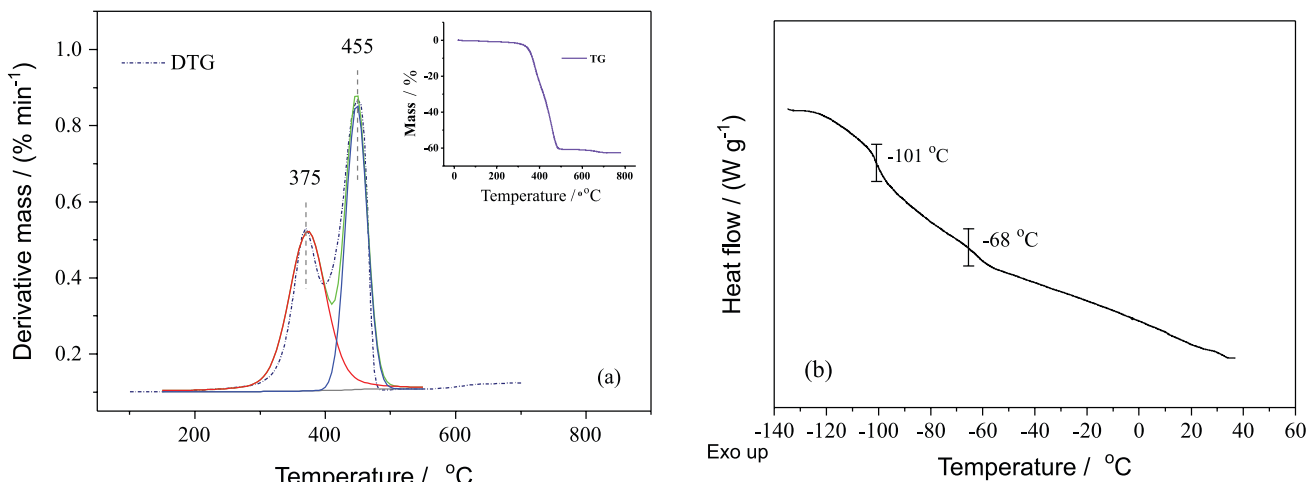


Figure 2. Thermal analysis of the vulcanizate. (a) TG-DTG and (b) DSC.

shown in Table 1.¹⁷ It is worth noting that the levels of pressure were all greater than the critical value of ethanol.

Table 1. Factors and levels of experimental design

Factor code	Minimum level -1	Central level 0	Maximum level +1
Temperature / °C	200	230	260
Pressure / MPa	8	10	12
time / min	30	75	120

The experimental design and its corresponding results are shown in Table 2. Three central point experiments were used to verify the accuracy, reproducibility and experimental error of regression model. The results of variant analyses are shown in Table 3. The influence of factors interaction on devulcanization was neglected because the main effect model agreed well with the experimental results. The probability value (*p*-value) of temperature was lower than 0.05 for SF and CD, and therefore temperature was

Table 2. Full factorial design and experimental responses (SF and CD)

Experiment	Temperature / °C	Pressure / MPa	time / min	SF ^a / %	CD ^b / (mmol cm ⁻³)
E1	260	8	120	36.96	0.133
E2	260	8	30	28.23	0.117
E3	200	8	30	3.53	0.141
E4	230	10	75	16.18	0.127
E5	260	12	30	33.00	0.069
E6	260	12	120	44.21	0.074
E7	230	10	75	18.17	0.132
E8	200	8	120	3.78	0.150
Center 1	200	12	120	4.22	0.141
Center 2	200	12	30	3.75	0.121
Center 3	230	10	75	15.38	0.122

^aSol fraction; ^bcrosslinking density of gel. E: experiment.

Table 3. Analysis and statistical factors of regression model of SF and CD

Variable	DF ^a	SF			CD		
		SS ^b	F ^c	<i>p</i> -Value ^d	SS ^b	F ^c	<i>p</i> -Value ^d
Model	3	2090.05	55.09	< 0.0001	0.00591	13.38	0.0028
T	1	2017.36	159.52	< 0.0001	0.00324	21.99	0.0022
P	1	19.84	1.57	0.2506	0.00235	15.93	0.0052
t	1	52.84	4.18	0.0802	0.00328	2.23	0.1794
Residual	7	88.52			0.00103		
Lack of fit	5	84.37	8.12	0.1132	0.000986	8.86	0.1045
Pure error	2	4.16			0.0000445		
Corrected total	10	2178.57			0.00694		

^aDegrees of freedom; ^bsum of square deviations from the overall mean; ^cF distribution ratio; ^dprobability value. SF: sol fraction; CD: crosslinking density of gel; T: temperature; P: pressure; t: time.

proven to be the only significant factor. Pressure and time had no appreciable effect on devulcanization. The normal probability plots of the residuals for responses are shown in Figure 3; no abnormal experimental points were observed.

For further analysis, the influence of temperature on devulcanization was studied (Figure 4). It could be observed that, in a low temperature region (from 210 to 240 °C), the SF increased gradually from 7 to 13% (Figure 4a). As the reaction temperature changed from 240 to 270 °C, the SF rose dramatically from 13 to 55%, however, it turned to depend weakly on the temperature around 56% in a higher temperature region (above 270 °C). There was an obvious difference in the increase rate of SF below and above 240 °C which was around the critical temperature of ethanol. This result showed that the devulcanizing process was accelerated when the ethanol entered into the supercritical state. The CDs at different temperatures are also shown in Figure 4b, and it can be seen that the CDs of devulcanized samples are all less than that of vulcanizate.

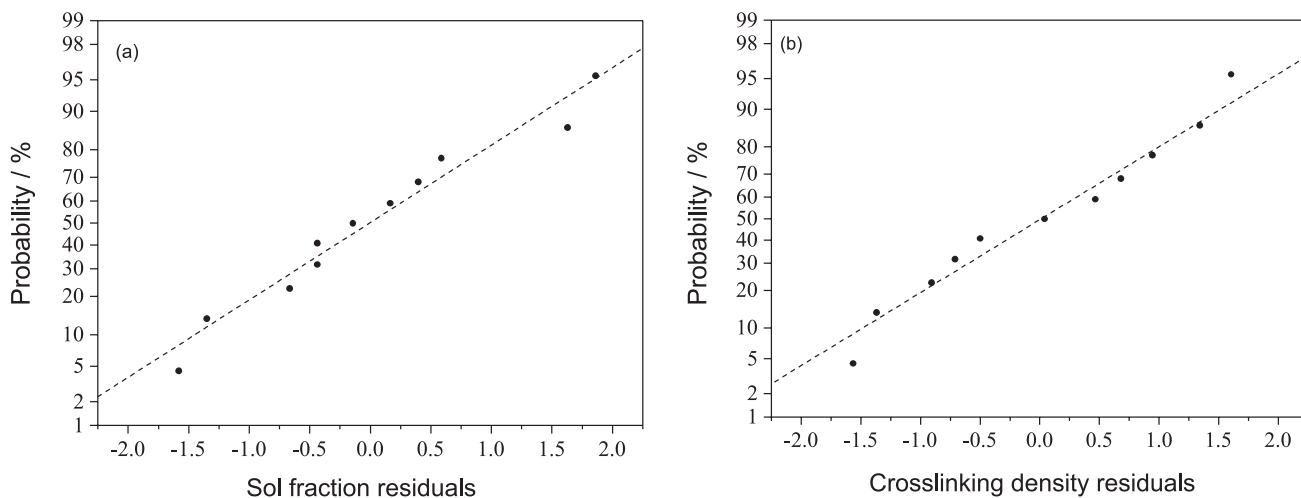


Figure 3. Normal probability plot of residuals for (a) SF and (b) CD of gel.

This implied that the supercritical ethanol permeated into the inside of vulcanizate and broke the crosslinked network. In supercritical ethanol, the CD of devulcanized samples decreased as the temperature goes up, and then, showed a slight increase when the temperature reached 270 °C. It indicated that the reaction of crosslinking by the unsaturated butadiene units occurred again at a higher temperature.²⁹ In the meantime, the molecular weight of sols decreased and the PDI increased as the reaction temperature rose (Table 4), revealing that some main chains were broken at the higher temperature.

In the previous study,¹⁹ we found that devulcanizing reagent was necessary for the devulcanizing reaction in ScCO₂. In order to analyze the influence of temperature with devulcanizing reagent in supercritical ethanol, DPDSs in various concentrations were added in the reaction. The result is presented in Figure 5. It can be seen that the SF increased as the temperature rose at a certain

Table 4. Characters of sol at various reaction temperatures

Temperature / °C	Mn ^a / (g mol ⁻¹)	Mw ^b / (g mol ⁻¹)	PDI ^c
230	64659	88937	1.375
250	40362	67903	1.682
270	37556	74583	1.986

^aNumber average molecular weight; ^bweight average molecular weight; ^cpolymer dispersity index.

concentration of DPDS. In the case of a relatively low temperature such as 250 °C, the SF increased obviously from 25 to 51% with an increase of DPDS from 0 to 12 g L⁻¹. However, the influence of DPDS was not obvious when the temperature reached a relatively high plateau such as 270 °C. This result demonstrated that the devulcanizing reagent could accelerate the reaction rate at a relatively low temperature, but the effect was waning with the increase in temperature.

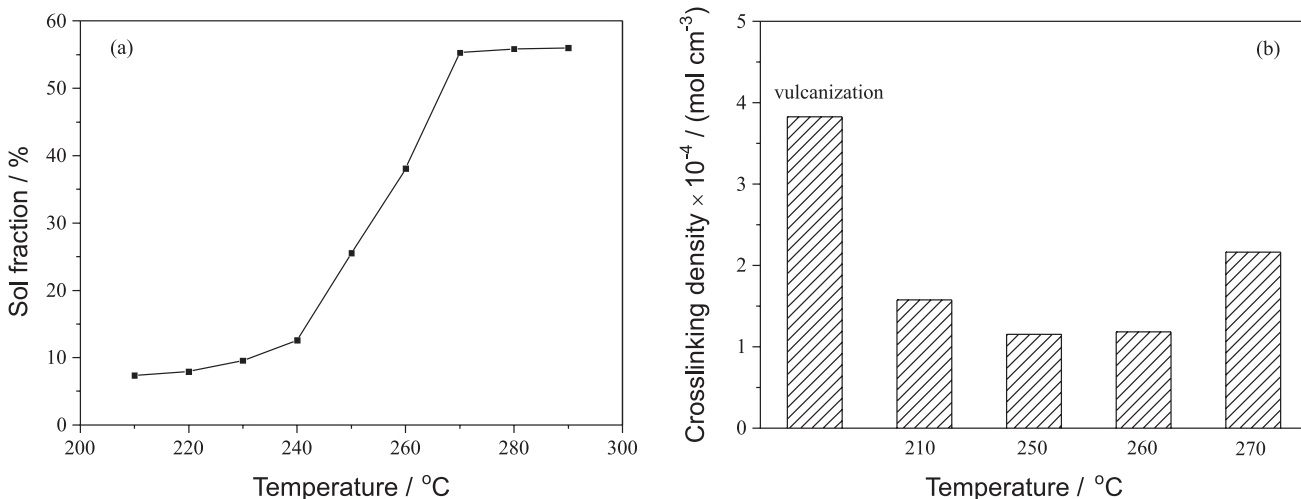


Figure 4. Effects of reactive temperature on (a) sols and (b) CD of gels. Reaction conditions: 8 MPa pressure, 60 min time.

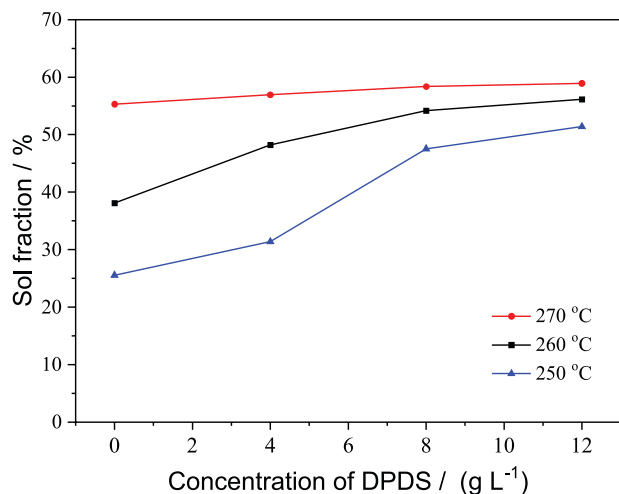
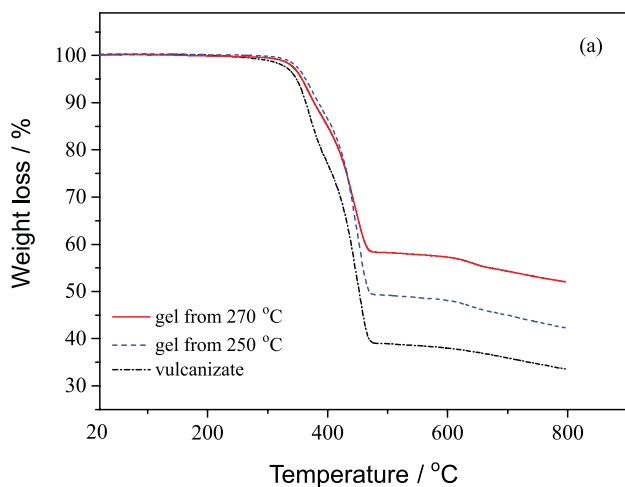


Figure 5. Effect of DPDS on devulcanization in supercritical ethanol. Reaction conditions: 8 MPa pressure, 60 min time.

In order to explore the effect of temperature on devulcanization mechanism in supercritical ethanol, the devulcanized samples compared with the vulcanizate were characterized and analyzed. All the devulcanized samples were obtained under the conditions of 8 MPa and 60 min.

Thermal analysis

The results of thermal analysis of vulcanizate and devulcanized gels are shown in Figure 6. The TG curve shapes between 300 and 500 °C are almost the same (Figure 6a), indicating that the thermostability of devulcanized gels was similar to the vulcanizate. Moreover, the weight loss of devulcanized gels from 300 to 500 °C decreased gradually with an increase in the reaction temperature. This can be attributed to the decrease of the polymer fraction in gels, which was in good conformity with the result in Figure 4a. Notably, the weight loss curve



of devulcanized gel shows a deviation to higher temperature owing to the proportional change of NR and BR in gels.³⁰ In DTG, the peak intensity at 375 °C is weakened gradually with an increase in temperature (Figure 6b). Through fitting analysis, the proportion of NR and BR in gels is shown in Table 5. It can be found that, after the devulcanization at 250 and 270 °C, the percentages of NR (39.2 and 33.5%, respectively) in gels reduced significantly in comparison with that in vulcanizate (53.6%). Correspondingly, the ratios of NR/BR among the three samples, vulcanizate and devulcanized gels at 250 and 270 °C, tended to decrease. These results implied that the devulcanizing reaction rate of NR was greater than that of BR in supercritical ethanol.

Table 5. Mass percentage of NR and BR through the DTG analysis

Sample	Peak position / °C		Proportion / %		NR/BR ratio
	NR	BR	NR	BR	
Vulcanizate	375	455	53.6	46.4	1.15
Gel at 250 °C	375	455	39.2	60.8	0.64
Gel at 270 °C	375	455	33.5	66.5	0.50

NR: natural rubber; BR: butadiene rubber.

Results of DSC analysis of vulcanizate and devulcanized gels are shown in Figure 7. Clearly, there were two glass transitions in the DSC curve of vulcanizate (in Figure 7a). However, the glass transition at -65 °C, according to NR, was not obvious in devulcanized gels (Figures 7b and 7c). This change implied that the most NR had been devulcanized and formed the sol in samples b and c, whose only glass transitions at -94 and -74 °C, respectively, were due to BR. In general, the mobility of polymer chains was strengthened with a decrease in CD, which would result in a lower glass transition temperature. However, Levin *et al.*³¹ found that the mono-, di- and polysulfur

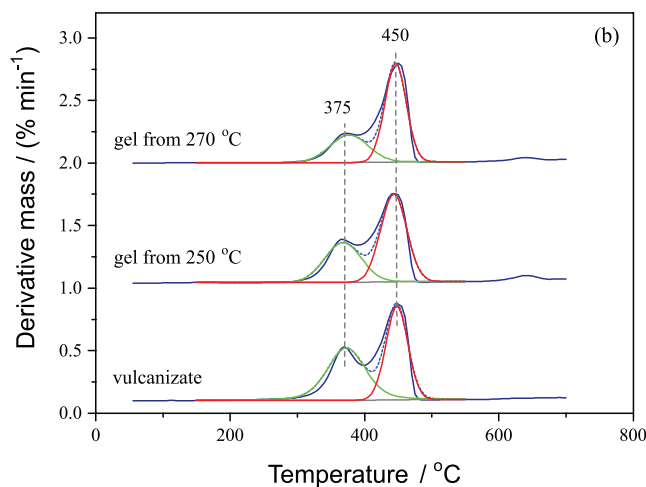


Figure 6. Thermal analysis of vulcanizate and devulcanized gels. (a) TG and (b) DTG.

bonds would rearrange into cyclic sulfur structures at a higher temperature, which might decrease the molecular mobility in polymer and result in a significant increase of glass transition temperature. Based on this explanation, the significantly increase in the glass transition temperature of gels (b and c) whose main polymer component was BR should be due to the rearrangement of bonds.

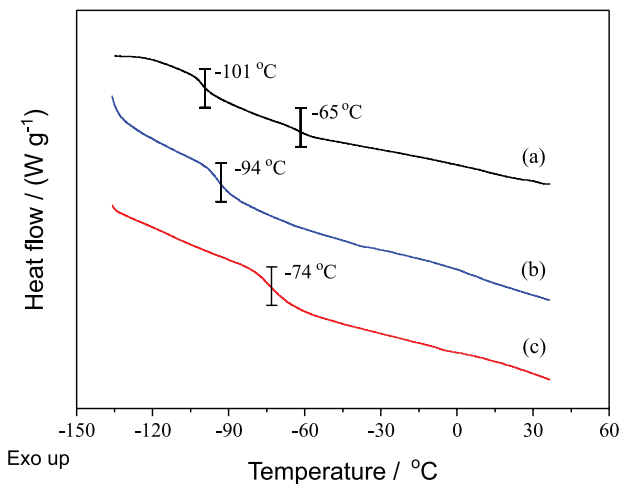


Figure 7. DSC analysis of vulcanizate and devulcanized gels. (a) Vulcanizate; (b) gel from devulcanization at 250 °C and (c) gel from devulcanization at 270 °C.

Structure analysis of sols at various reaction temperatures

The FTIR analysis of sols at different reaction temperatures is shown in Figure 8. The characteristic bands around 500, 570 and 730 cm^{-1} refer to the presence of S–S and C–S groups.^{32,33} It can be observed that the band of S–S (500 cm^{-1}) almost disappeared at different temperatures, confirming the breaking of a large proportion of the crosslink bonds, i.e., the success of the devulcanization treatment.³⁴

The bands located at 835 and 1372 cm^{-1} , attributed to C=CH stretching vibration and the $-\text{CH}_3$ stretching

vibration, respectively, confirm the presence of NR in the sol.^{35,36} The intensity of these two bands does not appreciably change with increment of treatment temperature, meaning that the C=C double bonds were not destroyed notably by the temperature rise. Therefore, a relatively complete polymer structure was maintained in sol.

The characteristic band at 970 cm^{-1} (the out-of-plane bending vibrations of C–H of *trans*-CH=CH-) corresponds to butadiene in BR.^{26,30,37} It is worth noting that, compared with other characteristic bands, the intensity of this band decreased with an increase in the reaction temperature, implying that the proportion of BR in the sol was reduced. It also proved that more NR was devulcanized to form the sol at higher temperature. This result agrees well with the change of NR/BR ratio in Table 5.

The bands between 1800 and 1500 cm^{-1} , corresponding to the C=C and C=O bonds, show that few changes take place at different temperatures.^{8,34} This result is in agreement with the analysis above, indicating that the supercritical ethanol could make the main chain structure maintain integrality.

Band of C–H bond located in the range from 3100 to 2750 cm^{-1} does not show differences at different temperatures,^{38,39} suggesting that the hydrocarbon backbone of sidewall rubber was not destroyed by higher temperature in supercritical ethanol.

In the general process of devulcanization reaction, there is an inverse relationship between the SF and CD of gel. The CD of the vulcanized rubber gradually declines until the polymer chain breaks away from the crosslinking network, and a sol is formed. However, the SF and CD of gel, here, increase simultaneously with the rising of temperature as shown in Figure 4. According to the analysis of DTG and FTIR, the main composition of sols is NR. Correspondingly, the main composition of gels is BR. The

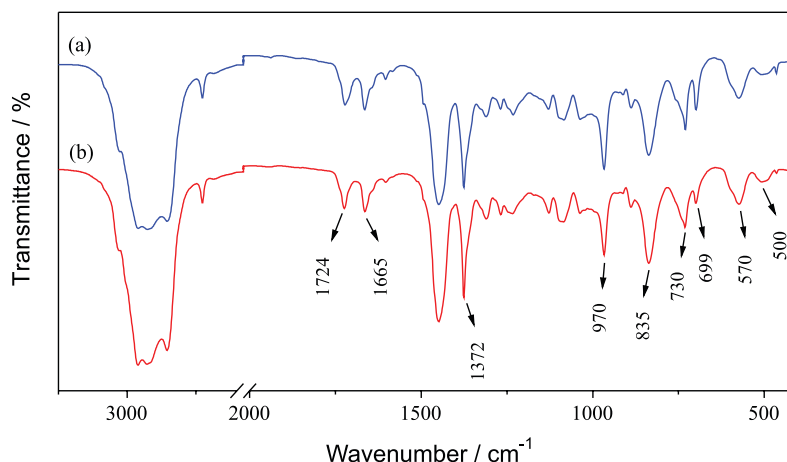


Figure 8. FTIR of sols at different reaction temperatures. (a) Sol from devulcanization at 250 °C and (b) sol from devulcanization at 270 °C.

high temperature causes more devulcanization of NR, but the increase in the CD of BR.

Model analysis of crosslinked network scission

Horikx's theory⁴⁰ was used to determine the breakage mechanism of crosslinked network during devulcanization. Figure 9 presents SF from devulcanizing sidewall rubber as a function of the relative decrease in CD at different temperatures. The solid curve corresponds to the devulcanizing process in which only crosslink bonds scission took place, and the dashed curve corresponds to the devulcanizing process in which only main chains scission occurred. When the temperature was 210 °C, the analysis result fell near the solid curve, indicating that the formation of sol was mainly caused by crosslink bonds scission. The result was close to the dashed curve as the temperature reached 250 °C, implying that the breakage of main chain gradually increased with the rising of temperature. When the temperature reached 260 or 270 °C, the results all fell above the dashed curve, indicating that the devulcanization was mainly caused by main chains scission. Obviously, main chains scission gradually predominated with the rise of temperature during devulcanization. It means that the increase of SF was mainly dependent on the breakage of main chains.

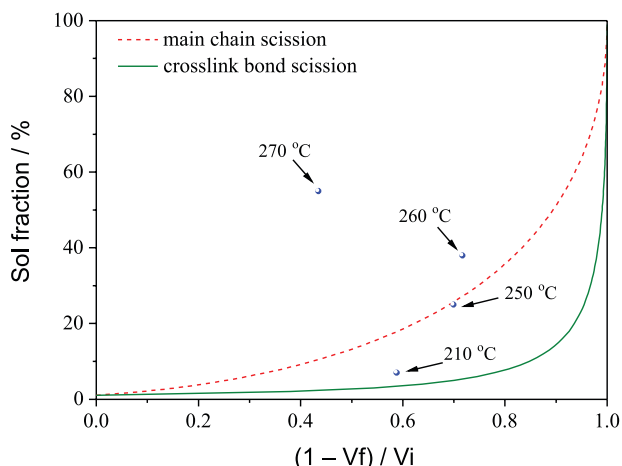


Figure 9. SF of devulcanized sidewall rubber against relative decrease in CD at different reaction temperatures.

Devulcanizing reaction rate

The reaction kinetics was used to further analyze the supercritical ethanol devulcanizing process. The objective function was gel concentration, and the calculation method is shown as equation 4. The initial gel concentration was 100%.

$$\text{Gel concentration (mg cm}^{-3}\text{)} = \frac{W_d \times w - W_s}{\text{Reactor volume (cm}^3\text{)}} \quad (4)$$

It was assumed that the devulcanizing process followed the first-order kinetics.^{14,20} The experimental results are shown in Figure 10. The first-order kinetics equation is shown as

$$-\frac{dC_{\text{gel}}}{dt} = kC_{\text{gel}} \quad (5)$$

where C_{gel} is the gel concentration, t is the reaction time and k is the reaction rate constant.

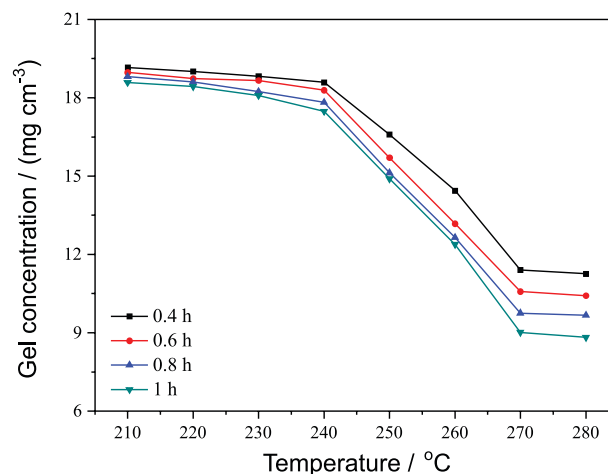


Figure 10. Experimental results for reaction kinetics.

The following equations 6 and 7 can be obtained by integrating equation 5 from t_0 to t (starting time and reaction time, respectively). In this study, the reaction time during the elevation of temperature was ignored, and the integral was performed only within the isothermal period, thus the starting time t_0 was 0.

$$-\int_{C_{0,\text{gel}}}^{C_{\text{gel}}} \frac{dC_{\text{gel}}}{C_{\text{gel}}} = k \int_{t_0}^t dt \quad (6)$$

$$-\ln C_{\text{gel}} = kt - \ln C_{\text{gel},0} \quad (7)$$

where $C_{\text{gel},0}$ is gel concentration at $t = t_0$. The results of plots in Figure 11a were obtained by equation 7. Further, the lines were calculated by the least-squared method and fitted well the measured variation of the gel concentration. The k value was calculated by the slope of lines as presented in Figure 11b. It can be observed that the k rose rapidly as the temperature was above 240 °C. This result further demonstrated that supercritical ethanol played an important role in promoting devulcanization.

Discussion

In general, the crosslinked network is formed by mono-,

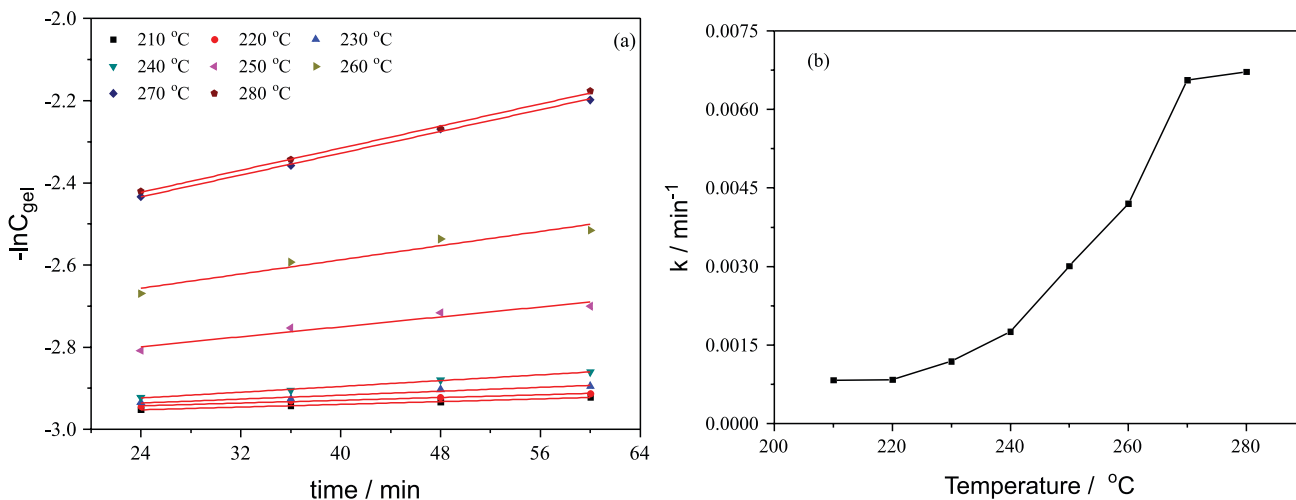


Figure 11. Reaction rate constant. (a) $-\ln C_{gel}$ vs. time and (b) k value vs. temperature.

di- and polysulfur bonds.^{41,42} In this study, with the increase in temperature, not only the crosslink bonds but also some main chains were broken in supercritical ethanol process (Figure 9), which all had great influence on SF.

In Figure 4, there can be three regions of temperature: a low temperature region below 240 °C in subcritical state, a medium temperature region from 240 to 270 °C and a high temperature region above 270 °C. In the low temperature region, the breakage of crosslinked bonds predominated the devulcanizing process, but reaction rate was relatively slow.

As the temperature reached the medium temperature region, the ethanol transferred to supercritical state. On one hand, the number of radicals increased due to the promotion on dynamics. The radicals had an instinctive tendency to couple and then form macromolecular structure fragments, which was the main compositions of sols.⁴³ On the other hand, the reaction activity of supercritical ethanol enhanced. As a result, the crosslinked structure broke rapidly and the sols increased sharply, with the unexpected occurrence of the main chain scission. Interestingly, due to the different structure properties between NR and BR, NR was easier to be devulcanized than BR as can be seen from the results of DTG and FTIR.

When the temperature further rose above 270 °C into the high temperature region, the degradation rate of crosslinked structure was fast enough to ignore the effect of DPDS. Meanwhile, the main chains continued breaking, which predominated in the devulcanizing process. In addition, the cyclic sulfide structure transformed from the rearrangement of crosslinking bonds grafted onto the main chains, which resulted in the increase of glass transition temperature (as shown in Figure 7).

Based on the presented discussion, the proposed mechanisms for both NR and BR are presented in

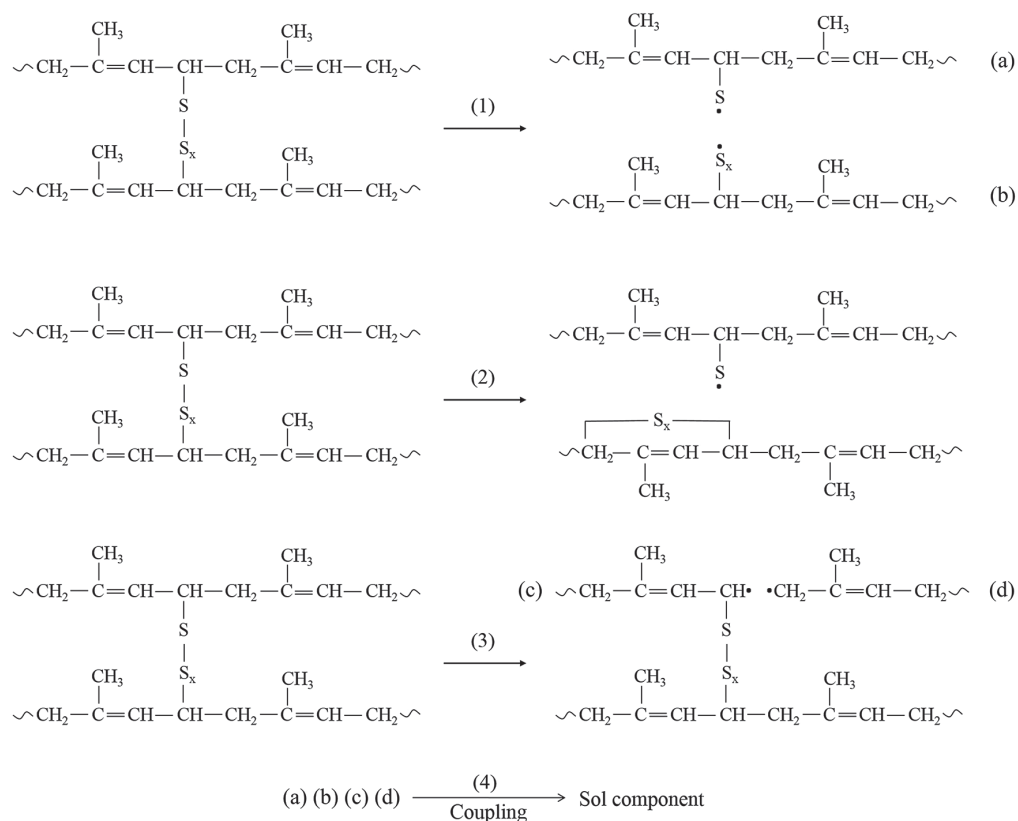
Schemes 1 and 2. According to the result of Figure 4b, the unsaturated butadiene units in BR would transform to the crosslink structure at a high temperature as a result of devulcanization by ethanol. The possible reaction process of increase in CD is also proposed in Scheme 3.

Conclusions

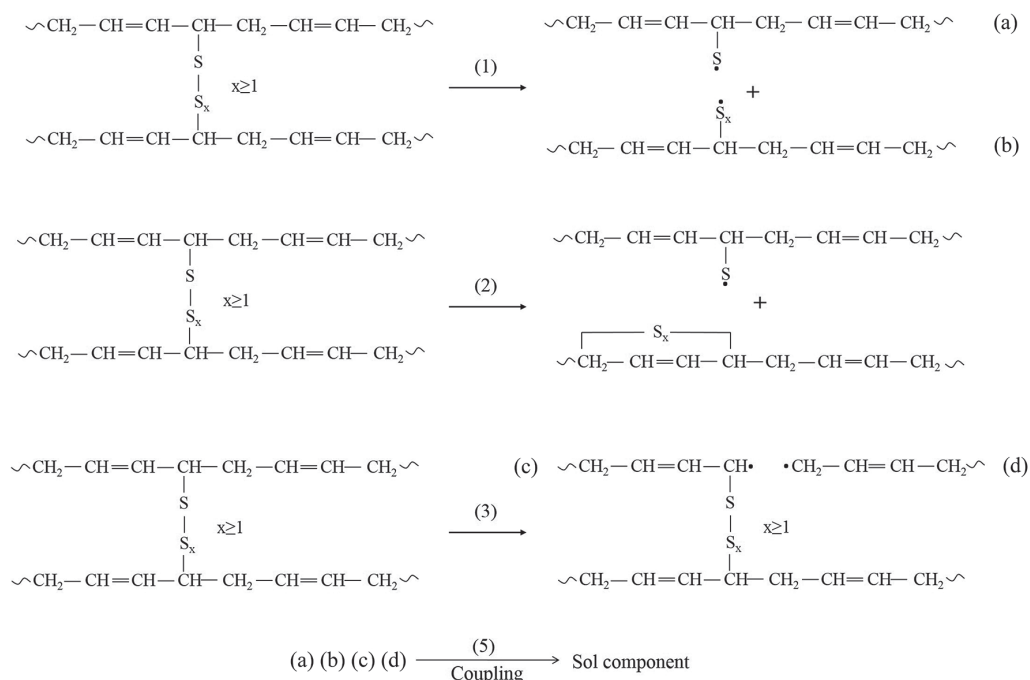
The effect of temperature on devulcanization for sidewall rubber in supercritical ethanol was studied. A full factorial experimental design was used to define the experimental conditions. Reaction temperature was the most important factor in the devulcanization process. Reaction pressure above critical value and reaction time exert no significant influence. When the reaction temperature changed from 240 to 270 °C, the SF rose dramatically from 13 to 55%. Meanwhile, the CD of devulcanized gel drastically reduced comparing with that of vulcanizate. The use of devulcanizing reagent did not significantly affect the SF at a high temperature of 270 °C.

Devulcanizing reaction rate of NR was faster than that of BR in supercritical ethanol. The mass percentage of NR/BR in gels continued to decrease with the rise in reaction temperature as revealed by DTG analysis, which was in agreement with the results of DSC and FTIR.

The scission of the crosslink bonds predominated in devulcanization process at lower temperatures and the main chains began to be broken when the temperature reached 250 °C, which has been proved by the mode analysis of the crosslinked network scission. The glass transition temperature of the gels increased with the rise in reaction temperature because of the reduction in the mobility of main chains, which was caused by the cyclic sulfur structure. The kinetic constant k rose significantly



Scheme 1. Simplified mechanism proposed for the formation of sol during devulcanization of NR by supercritical ethanol.

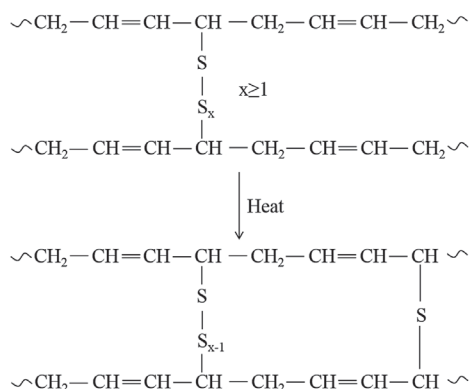


Scheme 2. Simplified mechanism proposed for the formation of sol during devulcanization of BR by supercritical ethanol.

when the temperature was above 240 °C, which indicated that temperature had the most notable promoting effect in the supercritical ethanol devulcanizing process.

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Scheme 3. The crosslinking reaction caused by unsaturated butadiene units at high temperature.

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